

Time Allowed: 45 minutes

Gases and Pressure

SECTION I

Answer the following questions briefly as far as possible.

1. A gas is contained in a cylinder. For a constant mass, describe how it can be made to
 - (a) expand the pressure is
 - (i) doubled
 - (ii) a third its value
 - (iii) tripled
 - (iv) sixfold

2. Rank the following gases in increasing order

- | | |
|------------|-----------|
| (a) 1 mole | (d) 10 g |
| (b) 1 gram | (e) 100 g |

3. Explain how to measure the partial pressure of a dry gas that is collected over water when the total pressure is atmospheric pressure.

4. Explain how the amount of gas varies with the rate of effusion at equal rates.

5. a. Sketch the curve for water vapor's partial pressure versus temperature in the gaseous state.

Temp. (°C)	Pressure (kPa)
0	0.6
5	0.7
10	0.9
15	1.2
20	1.6
25	2.3



- b. The above graph is for water vapor. Sketch
 - i. Another gas that would provide the above curve
 - ii. Another gas that would provide the above curve

Chapter 11 Review Gases Section 1 Answer Key

G Psacharopoulos



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