

CHEMISTRY ACTIVITY
Chapter 13 Section 3: Liquids

Name _____ Period _____ Date _____

1. Describe the type of state according to the kinetic molecular theory.

Answer: In the liquid state, particles have enough energy to move past one another, but not enough to escape completely from the liquid.

2. List the properties of liquids.

Answer: High density, definite mass, indefinite shape, no fixed volume, surface tension, evaporation, boiling, condensation, viscosity, and expansion on cooling.

3. How does the kinetic molecular theory explain the following properties of liquids?

- a. Relatively high density

Answer: The particles are close together - they are only not in fixed positions as their molecules are.

- b. Ability to diffuse

Answer: Particles in a liquid are free to move about in random ways (in all directions).

- c. Ability to evaporate

Answer: In a liquid, all molecules have a range of speeds. Some will have high, fast enough to break other, attractive forces and escape into the air.

- d. Change in evaporation and expansion

Answer: Evaporation is the process in which a liquid's surface molecules escape to gas. Evaporation is the only process in which the liquid's mass decreases. The rate of evaporation depends on temperature, surface area, and humidity.

Chapter 10 Section 2 Chemistry

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