

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the parentheses.

- 1.** mole: per centimole/mole
- 2.** empirical formula: empirical formula
- 3.** molar mass: molar mass (molar weight) and molar mass (M_r)
- 4.** 10^{-9} : nanogram
- 5.** a description of the atoms present in a chemical formula: composition
- 6.** the lowest whole-number ratio of the atoms of the elements in a compound: empirical formula
- 7.** formula structure: the way atoms are put in a 10^{-23} unit
- 8.** the smallest practical amount of a compound: molar mass
- 9.** the total mass of 6.02×10^{23} representative particles of a substance: molar mass
- 10.** 10^{-9} particles: nanogram
- 11.** the smallest mass of a compound that will have mass of one mole of a compound: 6.02×10^{23} representative particles (R.P.)

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Completion Use the completion method to check your knowledge of the terms and facts concerning the concepts involved with a certain type of substance composition. Some ideas are given as follows.

Knowing what mole counting means, it may be defined as a single quantity called the mole. Representative particles in a mole of a substance is Avogadro's number.

In the case of a mole of a compound, atoms are put together in formula units of the substance of the compound. When you substitute the unit given for atoms (the value of the molar mass) of the compound, there are 6.02×10^{23} representative particles in a mole of a substance.

True or False

10.1 A mole mass is measured in grams.

10.2 A mole of particles contains 6.02×10^{23} atoms.

10.3 The representative particles of a mole compound is the formula: **True**

10.4 A mole of 10^{-23} compound of one gram of sodium and two atoms of chlorine.

10.5 The mole mass of nitrogen gas is 28.02 g/mole .

Problems

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A Loxley



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