

# 12.1

# CHEMICAL REACTIONS

## Section Review

### Objectives

1. Identify substances using a chemical change.
2. Identify and predict chemical reactions using the laws of conservation.
3. Apply the law of conservation of mass to chemical reactions.

### Vocabulary

- |                             |                    |                                       |
|-----------------------------|--------------------|---------------------------------------|
| 1. <b>chemical change</b>   | 2. <b>reactant</b> | 3. <b>product</b>                     |
| 4. <b>chemical reaction</b> | 5. <b>product</b>  | 6. <b>law of conservation of mass</b> |

## Part A Completion

Write the appropriate words in the blanks. Use the key words in the box to help you. Write the number of the blank in the space provided.

- |  |  |
|--|--|
| 1. _____ change involves a change in _____.  | 1. <b>reactant</b>                     |
| 2. _____ change involves a change in _____.  | 2. <b>product</b>                      |
| 3. The _____ of a _____ change is the _____. | 3. <b>product</b>                      |
| 4. _____ is the _____ of a _____.            | 4. <b>reactant</b>                     |
| 5. _____ change involves a _____.            | 5. <b>chemical reaction</b>            |
| 6. _____ change involves a _____.            | 6. <b>law of conservation of mass</b>  |
| 7. _____ change involves a _____.            | 7. <b>product</b>                      |
| 8. _____ change involves a _____.            | 8. <b>reactant</b>                     |
| 9. _____ change involves a _____.            | 9. <b>chemical reaction</b>            |
| 10. _____ change involves a _____.           | 10. <b>law of conservation of mass</b> |

## Part B True-False

Write true or false in the space provided. Use the key words in the box to help you.

1. \_\_\_\_\_ change involves a change in \_\_\_\_\_.
2. \_\_\_\_\_ change involves a change in \_\_\_\_\_.
3. \_\_\_\_\_ change involves a change in \_\_\_\_\_.

# 2neutralization Reactions Section Review

**Daniela Niemeyer**



## **2neutralization Reactions Section Review:**

## **2neutralization Reactions Section Review** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**2neutralization Reactions Section Review**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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 l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients  $\text{Cu} + \text{MnO}_4^{-} + 4\text{H}^{+}$  is

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}^{2+} + \text{O}_7^{2-} + 2\text{Fe}^{3+} + 3\text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 + \text{H}_2\text{O}_2 + 2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{Au} + \text{CN}^- + 2\text{e}^-$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 + \text{H}_2\text{O}_2 + 2\text{OH}^-$  3 the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh 4 in basic solution created by jay

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web bonus problem here is a balanced net ionic equation  $3\text{Br}^- + 2\text{OH}^- + 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  o the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side

is 0 oxidation number of nickel is 2 now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + \text{Fe}^{2+} + \text{Fe}^{3+} + 2\text{e}^-$  equalize the electrons  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 6\text{Fe}^{2+} + 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 + \text{SO}_4^{2-} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 9\text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{TeO}_4$  11  $\text{PbO}_2 + \text{IO}_2 + \text{I}_2 + \text{Pb}^{2+}$  12  $\text{AsO}_3 + \text{NO}_3^- + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 + \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  b  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{C}_6\text{H}_6\text{O}_6$  c  $\text{H}_2\text{O}$  and  $\text{C}_6\text{H}_8\text{O}_6$  d  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{C}_6\text{H}_6\text{O}_6$

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