AP Biology Study Guide 1

Water Properties:

Cohesion- Hydrogen bonding between LIKE molecules
(ex. water with water)
Adhesion- Hydrogen bonding between UNLIKE molecules
(ex. water with a leaf)

Capillary Action- the movement of a liquid through a narrow space as a result of the forces of cohesion, adhesion, and surface tension.
(ex. Moving ground water from wet areas of soil to dry ones.)

Universal Solvent- Water is referred to as a universal solvent because it dissolves more substances than any other liquid.

Hydration Shell- the sphere of water molecules around each dissolved ion.

Moderation of Temperature- water moderates air temperature by absorbing heat from the air that is warmer and releasing stored heat to the air that is cooler.

Density- water is LESS dense as a SOLID than a LIQUID
(this is because of water's hydrogen bonds)

Polarity- water is a polar molecule because of the unequal distribution of electrons.

High Specific Heat- water has an unusually HIGH specific heat, this is how water is able to maintain its homeostasis and keep the ocean habitable.

Macromolecules:

Carbohydrates:

Elemental Makeup: Carbon, Hydrogen and Oxygen

Monomer: Monosaccharides (ex. Glucose & Fructose)

Polysaccharides:
- Cellulose: Strengthens cell walls
- Starch: Stores glucose for energy (plants)
- Glycogen: Stores glucose for energy (animals)
- Chitin: Strengthens exoskeletons and fungal cell walls

Key Function: Fuel, Carbon source to be converted to other molecules or into polymers

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