

# Chemistry — Answer Sheet

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1) Which of the following is a strong oxidising agent? (At. No. Mn = 25, Zn = 30, Cr = 24, Sc = 21)

Options: a)  $Mn^{4+}$  b)  $Zn^{2+}$  c)  $Cr^{3+}$  d)  $Sc^{3+}$  Answer: a)  $Mn^{4+}$  (e.g.,  $MnO_2$  / Mn(IV) species are good oxidising agents because Mn(IV) can gain electrons to be reduced to lower, more stable oxidation states.)

2) Why are Zn, Cd and Hg non-transition elements? Answer: By definition a transition element has an incomplete d-subshell in one or more common oxidation states.  $Zn^{2+}$ ,  $Cd^{2+}$  and  $Hg^{2+}$  have d10 configurations (completely filled d-subshell) and therefore do not show typical transition-metal behaviour (variable oxidation states, prominent d-d transitions). So they are not transition elements.

3) Why is  $Cu^{2+}$  coloured while  $Zn^{2+}$  is colourless in aqueous solution? Answer:  $Cu^{2+}$  has a d9 electronic configuration so d-d transitions (or charge-transfer transitions) are possible → coloured.  $Zn^{2+}$  is d10, with no available d-d transitions → colourless in many complexes/aqueous solutions.

4) Out of zinc and tin, whose coating is better to protect iron from corrosion? Answer: Zinc. Zinc coating (galvanization) provides sacrificial protection because Zn is more electropositive than Fe

# Core Teaching Resources Chemistry Answer Sheet

**Terry C. Jones**



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