

Section 13.1

The Nature of Gases

OBJECTIVES:

- Describe the assumptions of the “kinetic theory” as it applies to gases.
- Interpret gas pressure in terms of kinetic theory.
- Interpret gas pressure in terms of kinetic theory.
- Define the relationship between Kelvin temperature and average kinetic energy.

- *Kinetic* refers to _____
- The energy an object has because of its motion is called _____
- The kinetic theory states that the tiny particles in all forms of matter are in _____

- *Three basic assumptions* of the kinetic theory as it applies to gases:

#1. Gas is _____ - usually molecules or atoms

- Small, hard spheres
- Insignificant volume; relatively far apart from each other
- No attraction or repulsion between particles

#2. Particles in a gas move rapidly in _____ motion

- Move in straight paths, changing direction only when colliding with one another or other objects
- Average speed of O₂ in air at 20 °C is an amazing 1700 km/h!

#3. Collisions are _____ - meaning kinetic energy is transferred without loss from one particle to another- the total kinetic energy remains constant

- Gas Pressure – defined as _____

- Due to:
 - a) _____
 - b) _____
- No particles present? Then there cannot be any collisions, and thus no pressure – called a _____

- Atmospheric pressure results from the collisions of air molecules with objects
 - Decreases as you climb a mountain because the air layer thins out as elevation increases

- _____ is the measuring device for atmospheric pressure, which is dependent upon weather & altitude

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