Name Key

1. Hydrogen, H_2, is used as rocket fuel. The hydrogen is burned in oxygen, O_2, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 is burned?

$$2\ H_2(g) + O_2(g) \rightarrow 2\ H_2O(g) \quad ; \Delta H = -484\ kJ$$

$$200.0\ g\ H_2 \left| \frac{1\ mol\ H_2}{2.016\ g\ H_2} \right. = 99.2\ mol\ H_2 \qquad 99.2\ mol\ H_2 \left| \frac{-484\ kJ}{2\ mol\ H_2} \right. = -24{,}006\ kJ$$

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.

$$CS_2(l) + 3\ O_2(g) \rightarrow CO_2(g) + 2\ SO2(g) \quad ; \Delta H = -1077\ kJ$$

What is ΔH for the following reaction?

reverse so change sign of ΔH & cut in ½ quantities

$$\tfrac{1}{2}\ CO_2(g) + SO_2(g) \rightarrow \tfrac{1}{2}\ CS_2(l) + \tfrac{3}{2}\ O_2(g)$$

$$\frac{+1077\ kJ}{2} = 538.5\ kJ$$

3. Following is the equation for the combustion of acetone:

$$2\ C_3H_6O(g) + 8\ O_2(g) \rightarrow 6\ H_2O(g) + 6\ CO_2(g) + 853.6\ kcal$$

Exothermic

How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6\ kcal}{2\ mol\ C_3H_6O} \left| \right. 4.5\ mol\ C_3H_6O = -1920.6\ kcal$$

4. The specific heat of iron is 0.110 cal / g·°C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

$$heat = q = m\,c\,\Delta T = (5.2\ g)\left(0.110\ \frac{cal}{g\cdot°C}\right)(225 - 25°C) = 114.4\ cal$$

endothermic

5. Hydrazine, N_2H_4, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.

$$(1) \quad N_2(g) + 2H_2(g) \rightarrow N_2H_4(l)$$

$$N_2H_4(l) + O_2(g) \rightarrow N_2(g) + 2H_2O(l) \quad ; \Delta H = -622.2\ kJ$$

$$H_2(g) + \tfrac{1}{2}\ O_2(g) \rightarrow H_2O(l) \quad ; \Delta H = -285.8\ kJ$$

Reverse

$$N_2(g) + 2H_2O(l) \rightarrow N_2H_4(l) + O_2(g) \qquad \Delta H = 622.2\ kJ$$

doubly quantity

$$2H_2(g) + O_2(g) \rightarrow 2H_2O(l) \qquad \Delta H = (2)(-285.8)\ kJ$$

$$N_2(g) + 2H_2(g) \rightarrow N_2H_4(l) \qquad \Delta H = +50.6\ kJ$$

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