

## Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of  $\text{C}_6\text{H}_{12}\text{O}_6$ .

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia ( $\text{NH}_3$ ) can be produced?

3) Assume that in the decomposition of potassium chlorate,  $\text{KClO}_3$ ,



80.5 moles of  $\text{O}_2$  form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin"  $\text{PtCl}_2(\text{NH}_3)_2$ , has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g  $\text{K}_2\text{PtCl}_4$ ?

B. How many moles of  $\text{NH}_3$  would be needed if 3.8 moles of  $\text{K}_2\text{PtCl}_4$  were present?

# Chapter 12 Stoichiometry Test Answer Key

**Thomas Evangelist**



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## **Table of Contents Chapter 12 Stoichiometry Test Answer Key**

1. Understanding the eBook Chapter 12 Stoichiometry Test Answer Key
  - The Rise of Digital Reading Chapter 12 Stoichiometry Test Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 12 Stoichiometry Test Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 12 Stoichiometry Test Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 12 Stoichiometry Test Answer Key
  - Personalized Recommendations
  - Chapter 12 Stoichiometry Test Answer Key User Reviews and Ratings
  - Chapter 12 Stoichiometry Test Answer Key and Bestseller Lists
5. Accessing Chapter 12 Stoichiometry Test Answer Key Free and Paid eBooks
  - Chapter 12 Stoichiometry Test Answer Key Public Domain eBooks
  - Chapter 12 Stoichiometry Test Answer Key eBook Subscription Services
  - Chapter 12 Stoichiometry Test Answer Key Budget-Friendly Options
6. Navigating Chapter 12 Stoichiometry Test Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 12 Stoichiometry Test Answer Key Compatibility with Devices
  - Chapter 12 Stoichiometry Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 12 Stoichiometry Test Answer Key
  - Highlighting and Note-Taking Chapter 12 Stoichiometry Test Answer Key
  - Interactive Elements Chapter 12 Stoichiometry Test Answer Key

8. Staying Engaged with Chapter 12 Stoichiometry Test Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 12 Stoichiometry Test Answer Key
9. Balancing eBooks and Physical Books Chapter 12 Stoichiometry Test Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 12 Stoichiometry Test Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 12 Stoichiometry Test Answer Key
  - Setting Reading Goals Chapter 12 Stoichiometry Test Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 12 Stoichiometry Test Answer Key
  - Fact-Checking eBook Content of Chapter 12 Stoichiometry Test Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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