

## Acids and Bases Worksheet

NAME \_\_\_\_\_

- Name the following compounds as acids and circle which are weak acids:  
 a)  $\text{HNO}_3$       b)  $\text{H}_2\text{SO}_4$       c)  $\text{HF}$       d)  $\text{H}_2\text{CO}_3$       e)  $\text{HC}_2\text{H}_3\text{O}_2$       f)  $\text{H}_3\text{PO}_4$
- Write the formulas for the following bases:  
 a) calcium hydroxide    b) aluminum hydroxide    c) ammonium hydroxide    d) lithium hydroxide
- Complete the following table:

	Acids	Bases
Definition		
Electrolyte in aqueous solution?		
pH		
Taste		
Litmus turns		
Which reacts with metals to form $\text{H}_2$ gas?		

### Conjugate acid-base pairs

- Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:  
 $\text{NH}_3(\text{g}) + \text{HOH}(\text{l}) \rightleftharpoons \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq})$
- Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:  
 $\text{HCO}_3^-(\text{aq}) + \text{HOH}(\text{l}) \rightleftharpoons \text{CO}_3^{2-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$
- An acid will form the conjugate \_\_\_\_\_ and a base will form the conjugate \_\_\_\_\_.

### Neutralization and Titration.

- Classify the following acids as monoprotic, diprotic or triprotic:  
 a)  $\text{H}_2\text{SO}_4$       b)  $\text{H}_3\text{PO}_4$       c)  $\text{HCl}$       d)  $\text{HC}_2\text{H}_3\text{O}_2$  or  $\text{CH}_3\text{COOH}$  (same thing)
- What is the concentration of an unknown  $\text{NaOH}$  solution if took an average of 15.4mL of 0.100N  $\text{HCl}$  solution to neutralize 10.0mL of the  $\text{NaOH}$  solution?
- From the given data, what is the molarity and the normality of the  $\text{HF}$  solution if the  $\text{KOH}$  solution was 0.50M?

	Acid, $\text{HF}$	Base, $\text{KOH}$
Initial Volume, mL	5.5	14.9
Final Volume, mL	25.9	24.4

**$\text{pH} = -\log[\text{H}_3\text{O}^+]$**  To find pH if given  $[\text{H}_3\text{O}^+]$ , enter the  $[\text{H}_3\text{O}^+]$ , press Log & then the +/- button  
 To find  $[\text{H}_3\text{O}^+]$  if given pH, enter negative pH, then 2<sup>nd</sup> function Log

- A solution has a  $[\text{H}_3\text{O}^+]$  of  $1 \times 10^{-7}$  M, what is its pH? \_\_\_\_\_ Is this an acid, base or is it neutral?

TURN THE PAPER OVER

- What is the pH of a solution if the  $[\text{H}^+] = 7.2 \times 10^{-9}\text{M}$ ? \_\_\_\_\_ Is this an acid or base?

# Chapter 1acids And Bases Answer Key

**LL Leslie**



## **Chapter 1acids And Bases Answer Key:**

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