

**14 • The Chemistry of Acids & Bases****STUDY QUESTIONS & PROBLEMS - KEY**

1. For the following aqueous equilibria, designate the Bronsted-Lowry conjugate acid-base pairs and establish the weaker side:



2. Write the name and formula for the conjugate bases of the following:

- a.  $\text{HNO}_2$        $\text{NO}_2^-$       nitrite  
 b.  $\text{H}_2\text{SO}_4$        $\text{HSO}_4^-$       bisulfate  
 c.  $\text{H}_2\text{PO}_4^-$        $\text{HPO}_4^{2-}$       dihydrogen phosphate  
 d.  $\text{HF}$        $\text{F}^-$       fluoride  
 e.  $\text{CH}_3\text{CO}_2\text{H}$        $\text{CH}_3\text{COO}^-$       acetate

3. Complete the Bronsted-Lowry equilibria, label the components acid or base, and pair up the conjugate acid-base pairs:



4. Is the monohydrogen phosphate ion  $\text{HPO}_4^{2-}$  amphiprotic? — YES  
 If so, write the formulas of its conjugate acid and its conjugate base. CA =  $\text{H}_2\text{PO}_4^-$  and CB =  $\text{PO}_4^{3-}$

# Chapter 17 The Chemistry Of Acids Bases Study Questions

**Jessica J Manson**



**Chapter 17 The Chemistry Of Acids Bases Study Questions:**

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