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> Practical Workbook answers

Chapter 1

Practical investigation 1.1

Getting started

Ice keeps a drink cold on a hot day as any heat that is absorbed by the drink is used to melt the ice rather than to increase the temperature of the liquid.

Recording data

- 1 Table: Units should be added to the table:
Time / minutes, Temperature / °C
- 2 Graph: axes – labelled correctly with time on *x*-axis and temperature on *y*-axis; scale – suitable for data / more the 50% of graph paper used; plot – all points plotted correctly (± 1 mm); line of best fit – smooth line passing through all points / points joined dot to dot
- 3 temperature, heating, intermolecular, molecules, melting, liquid, boiling, gas
- 4
 - a melting point – F
 - b boiling point – G
 - c solid – A
 - d liquid – C
 - e gas – E
- 5 Take readings more frequently / use a data logger to record data continuously. Repeat the experiment and calculate a mean temperature for melting and boiling points. Use a machine to stir the water.
- 6 The melting and boiling points would not be a fixed temperature but would occur over a range of temperatures / graph showing the same.

Practical investigation 1.2

Getting started

Students may suggest a variety of answers but some may already understand that energy changes accompany changes of state and may link this to temperatures remaining the same during state changes.

Recording data

- 1–5 Dependent on results.
- 6 Use of data logger / digital thermometer.
- 7 To allow greater control of the temperature / so that the stearic acid was not heated too strongly.
- 8 Water does not freeze until 0°C is reached and this is difficult to achieve in a laboratory.
- 9 Use an electrical water bath / use a hot plate / any reasonable suggestion.



Chapter 17 Chemistry Workbook Answers

E Durkheim



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