

Acid/Base Titration – pH Calculations

Strong acid titrated with strong base



- Issue is which is in excess, (XS), H^+ or OH^-
- **Pre-equivalence strategy:**
 1. H_2O^+ is in XS, not completely neutralized.
 2. $(M_a V_a - M_b V_b)/(V_a + V_b) = [\text{unreacted acid}]$
 3. $[\text{unreacted acid}] = [\text{H}^+]$
 4. Calculate pH
- **Equivalence** pH = 7.00
- **Post-equivalence strategy:**
 1. $(M_b V_b - M_a V_a)/(V_a + V_b) = [\text{XS base}]$
 2. pH = pK_a - pOH

Weak acid titrated with a strong base

– To calculate pH we must now look at the conjugate base as it hydrolyzes water molecules.

• Three different calculations:

1 pre-equivalence strategy:

a) determine moles /molarity of unreacted acid remaining:

initial moles acid - moles base added = moles acid left

$$(M_a V_a - M_b V_b)/(V_a + V_b) = [\text{unreacted acid}]$$

b) determine moles/molarity of conjugate base released;

moles base added = moles conjugate released

$$(M_b V_b)/(V_a + V_b) = [\text{conjugate}]$$

c) use K_a of the acid and calculate $[\text{H}^+]$ & pH.

Chapter 16 Review Acid Base Titration Ph Answers

Jürgen Polster, Heinrich Lachmann



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Chemistry** Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order

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