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Pages: 573 - 589	Chapter 16 - Waste Generation, Terrestrial Pollution, and Waste Disposal MODULE 51 GUIDE <i>"Only Humans Generate Waste"</i>		Due:
<i>Learning Goals:</i>	<ul style="list-style-type: none"> ● Explain why we generate waste and describe recent waste disposal trends ● Describe the content of the solid waste stream in the United States 		
<i>Key Term AD's (Flashcards):</i>			
<i>Waste</i>	<i>Municipal Solid Waste (MSW)</i>	<i>e-waste</i>	
<p>What is municipal solid waste, and why has human society produced increasing amounts of it over time?</p> <p>Municipal solid waste is waste that is collected by municipalities from households, prisons, schools, municipal buildings, hospitals, and other small businesses. There is a relationship between the amount of waste and population, for instance, as the population increases, the amount of municipal waste increases.</p> <p>What does the term "Throwaway Society" mean? Why is relevant to waste generation/management?</p> <p>Throwaway Society refers to how people use items and get rid of them immediately after. This is important to understand because this idea results in more waste generation. This leads to new solutions for waste management in the future.</p> <p>What are the main sources of waste generation in developed countries such as the United States?</p> <p>Currently the main sources of waste generation in the United States is left over food, plastic, and paper.</p> <p>When broken down by composition, what type of product makes up the largest fraction of MSW?</p> <p>Food, yard waste, and paper make up more than half of Municipal Solid Waste by its weight.</p> <p>What types of waste can be composted? What do these products have in common that allows this?</p> <p>Organic material is what makes up most of the things that are composted. Adding onto that, Food scraps as well as yard waste are the most common compostable items.</p> <p>What are the major challenges that E-Waste poses, and how has it been dealt with so far?</p> <p>E waste contains mercury, lead, cadmium, and other toxic metals. These toxic metals can leach into the groundwater around a landfill. To solve these problems, e-waste away from landfills.</p> <p>Why is waste reduction generally considered to be the most important & effective of the "3R's"?</p>			

Chapter 16 Guided Reading Science Urban Life

ML Yell



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{ag aq} + \text{cr}_2\text{o}_7^{2-} \text{aq} \rightarrow \text{ag}_2\text{cr}_2\text{o}_7 \text{s}$ label 4 2 3

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{na aq} + 2\text{cl aq} + \text{cu}_2\text{so}_4 \text{aq} \rightarrow 2\text{na aq} + \text{so}_4^{2-} \text{aq} + \text{cu}_2\text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{h aq} + \text{cl aq} + \text{na aq} + \text{hco}_3 \text{aq} + \text{na aq} + \text{cl aq} + \text{h}_2\text{o l} + \text{co}_2 \text{g}$ the spectator ions are cl and na net ionic equation $\text{h aq} + \text{cl aq} + \text{na aq} + \text{hco}_3 \text{aq} + \text{na aq} + \text{cl aq} + \text{h}_2\text{o l} + \text{co}_2 \text{g} \rightarrow \text{h aq} + \text{hco}_3 \text{aq} + \text{h}_2\text{o l} + \text{co}_2 \text{g}$ here is a video on spectator ions

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web the six no3 aq ions and the six na aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{ba}^{2+} \text{aq} + 2\text{po}_4^{3-} \text{aq} \rightarrow \text{ba}_3\text{po}_4 \text{s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{h}^+\text{aq} + \text{oh}^-\text{aq} \rightarrow \text{h}_2\text{ol}$ the spectator ions in koh^+aq and hno_3^-aq are k^+aq and no_3^-aq how to identify and list the

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web the ag^+aq and cl^-aq ions become agcl_s but the na^+aq ions and the no_3^-aq ions stay as na^+aq ions and no_3^-aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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