

17

WATER AND AQUEOUS SYSTEMS

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SECTION 17.1 LIQUID WATER AND ITS PROPERTIES (pages 475–478)

This section explains the hydrogen bonding that occurs in water molecules and how high surface tension and low vapor pressure can be explained in terms of hydrogen bonding.

► The Water Molecule (pages 475–476)

- What unique molecular compound is the foundation of all life on Earth?
water
- Approximately what fraction of Earth's surface is covered in water? 3/4
- Circle the letter next to each sentence that is true concerning water molecules.
 - Each O — H covalent bond in a water molecule is nonpolar.
 - b. In a water molecule, the less electronegative hydrogen atoms acquire a slight positive charge and the oxygen atom acquires a slight negative charge.
 - c. Because the water molecule has an H — O — H bond angle of 105°, the molecule as a whole is polar.
- The diagram below depicts a water molecule. Complete the labels showing the locations of the hydrogen atoms, the oxygen atom, and the regions of positive and negative charge.



- List five important properties of water that result from hydrogen bonding.
 - high surface tension
 - low vapor pressure
 - high specific heat capacity
 - high heat of vaporization
 - high boiling point

Chapter 15 Water And Aqueous Systems Worksheet Answers

Robert R. Redfield



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Table of Contents Chapter 15 Water And Aqueous Systems Worksheet Answers

1. Understanding the eBook Chapter 15 Water And Aqueous Systems Worksheet Answers
 - The Rise of Digital Reading Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 15 Water And Aqueous Systems Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Personalized Recommendations
 - Chapter 15 Water And Aqueous Systems Worksheet Answers User Reviews and Ratings
 - Chapter 15 Water And Aqueous Systems Worksheet Answers and Bestseller Lists

5. Accessing Chapter 15 Water And Aqueous Systems Worksheet Answers Free and Paid eBooks
 - Chapter 15 Water And Aqueous Systems Worksheet Answers Public Domain eBooks
 - Chapter 15 Water And Aqueous Systems Worksheet Answers eBook Subscription Services
 - Chapter 15 Water And Aqueous Systems Worksheet Answers Budget-Friendly Options
6. Navigating Chapter 15 Water And Aqueous Systems Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 15 Water And Aqueous Systems Worksheet Answers Compatibility with Devices
 - Chapter 15 Water And Aqueous Systems Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Highlighting and Note-Taking Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Interactive Elements Chapter 15 Water And Aqueous Systems Worksheet Answers
8. Staying Engaged with Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 15 Water And Aqueous Systems Worksheet Answers
9. Balancing eBooks and Physical Books Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 15 Water And Aqueous Systems Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Setting Reading Goals Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Fact-Checking eBook Content of Chapter 15 Water And Aqueous Systems Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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