

## 14

## THE BEHAVIOR OF GASES

## Chapter Test A

## A. Matching

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A	Column B
_____ 1. ideal gas constant ( $R$ )	a. The volume of a fixed mass of gas is directly proportional to the Kelvin temperature if the volume is kept constant.
_____ 2. Boyle's law	b. At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases.
_____ 3. Dalton's law of partial pressures	c. $8.31 \frac{\text{L}\cdot\text{kPa}}{\text{K}\cdot\text{mol}}$
_____ 4. ideal gas law	d. the contribution each gas in a mixture makes to the total pressure
_____ 5. combined gas law	e. A gas tends to move to an area of lower concentration until the concentration is uniform throughout.
_____ 6. Charles's law	f. $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$
_____ 7. diffusion	g. $P \times V = n \times R \times T$
_____ 8. partial pressure	h. For a given mass of gas at constant temperature, the volume of gas varies inversely with the pressure.

## B. Multiple Choice

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 9. As the temperature of a fixed volume of gas increases, the pressure will
- vary inversely.
  - decrease.
  - be unchanged.
  - increase.
- \_\_\_\_\_ 10. A breathing mixture used by deep-sea divers contains helium, oxygen, and carbon dioxide. What is the partial pressure of oxygen at 101.3 kPa total pressure if  $P_{\text{He}} = 84.0$  kPa and  $P_{\text{CO}_2} = 0.10$  kPa?
- 10.3 kPa
  - 17.2 kPa
  - 34.4 kPa
  - 185.4 kPa

# Chapter 14 The Behavior Of Gases Chapter Test A Answers

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