

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

States of Matter

Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

- The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
- The kinetic molecular theory makes the following assumptions:
 - In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
 - Between gas particles and the walls, there are no significant attractive or repulsive forces between gas particles.
 - Gas particles are in constant and random motion.
 - The collisions between gas particles are elastic, that is, no kinetic energy is lost.
- The kinetic energy of a particle is represented by the equation $KE = \frac{1}{2}mv^2$.
- Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- | | |
|--------------|--|
| <u>True</u> | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas. |
| <u>True</u> | 2. The random motion of gas particles causes a gas to expand until it fills its container. |
| <u>False</u> | 3. The density of a gas decreases as it is compressed. |
| <u>True</u> | 4. A gas can flow into a space occupied by another gas. |
| <u>True</u> | 5. The diffusion of a gas is caused by the random motion of the particles of the gas. |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles. |
| <u>True</u> | 7. During effusion, a gas escapes through a tiny opening into a vacuum. |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 13 79

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CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

Section 13.1 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

- Pressure is defined as force per unit
 - area
 - time
 - mass
 - volume
- What is an instrument designed to measure atmospheric pressure?
 - barometer
 - manometer
 - physoconimeter
 - densimeter
- The height of the liquid in a barometer is affected by all of the following EXCEPT the
 - altitude
 - density of the liquid in the column
 - atmospheric pressure
 - diameter of the column tube.
- The pressure of the gas in a manometer is directly related to which of the following
 - height of the mercury column in the closed-end arm
 - height of the mercury column in the open-end arm
 - $h + 76$
 - $h - 76$
- One atmosphere is equal to a pressure of
 - 76 mm Hg
 - 101.3 kPa
 - 147 psi
 - 760 mm
- The partial pressure of a gas depends on all of the following EXCEPT the
 - concentration of the gas
 - size of the container
 - density of the gas
 - temperature of the gas
- The pressure of a sample of air in a manometer is 101.3 kPa. What is the partial pressure of nitrogen (N_2) in the sample if the combined partial pressures of the other gases is 22.4 kPa?
 - 42.4 kPa
 - 78.9 kPa
 - 101.3 kPa
 - 123.7 kPa

Use the figure to answer the following questions.

- What instrument is illustrated in the figure? barometer
- Who invented this instrument? Evangelista Torricelli
- What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
- What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



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Chapter 13 States Of Matter Answers

United States



Chapter 13 States Of Matter Answers:

Superfluid States of Matter Boris V. Svistunov, Egor S. Babaev, Nikolay V. Prokof'ev, 2015-04-15 Covers the State of the Art in Superfluidity and Superconductivity Superfluid States of Matter addresses the phenomenon of superfluidity and superconductivity through an emergent topologically protected constant of motion and covers topics developed over the past 20 years The approach is based on the idea of separating universal classical field superfluidity and superconductivity *Science in Your World: Teacher resource masters* Jay K. Hackett, 1991 *The Science Teacher's Toolbox* Tara C. Dale, Mandi S. White, 2020-04-28 A winning educational formula of engaging lessons and powerful strategies for science teachers in numerous classroom settings The Teacher's Toolbox series is an innovative research based resource providing teachers with instructional strategies for students of all levels and abilities Each book in the collection focuses on a specific content area Clear concise guidance enables teachers to quickly integrate low prep high value lessons and strategies in their middle school and high school classrooms Every strategy follows a practical how to format established by the series editors The Science Teacher's Toolbox is a classroom tested resource offering hundreds of accessible student friendly lessons and strategies that can be implemented in a variety of educational settings Concise chapters fully explain the research basis necessary technology Next Generation Science Standards correlation and implementation of each lesson and strategy Favoring a hands on approach this book provides step by step instructions that help teachers to apply their new skills and knowledge in their classrooms immediately Lessons cover topics such as setting up labs conducting experiments using graphs analyzing data writing lab reports incorporating technology assessing student learning teaching all ability students and much more This book enables science teachers to Understand how each strategy works in the classroom and avoid common mistakes Promote culturally responsive classrooms Activate and enhance prior knowledge Bring fresh and engaging activities into the classroom and the science lab Written by respected authors and educators The Science Teacher's Toolbox Hundreds of Practical Ideas to Support Your Students is an invaluable aid for upper elementary middle school and high school science educators as well those in teacher education programs and staff development professionals **Answer Book to Accompany**

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United States Code Annotated United States,1927

United States Code United States,1954

Oregon Code, 1930 Oregon,1930

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