

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups rasp and blueberries, and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, KClO_3 . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{22} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(III) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 15.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. (Write a balanced equation first.)

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of nitrogen dioxide. Assume all gases are at the same conditions of temperature and pressure.



Chapter 12 Stoichiometry Section 12 Answer Key

Darrell D. Ebbing



Chapter 12 Stoichiometry Section 12 Answer Key:

Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's *Basic Concepts of Chemistry* provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. WileyPLUS sold separately from text.

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Chemical Problem-solving by Dimensional Analysis Arnold B. Loebel, 1978 **I/M&w/Tsts Intro Chem** Victor S. Krimsley, Darold E. Skerritt, Beverly B. Harrison, 1986 *McGraw-Hill's SAT Subject Test: Chemistry, 2ed* Thomas Evangelist, 2009-02-01 We want to help you score high on the SAT Chemistry test. We've put all of our proven expertise into McGraw-Hill's SAT Subject Test Chemistry to make sure you're fully prepared for this difficult exam. With this book, you'll get essential skill-building techniques and strategies created by a leading high school chemistry teacher. You'll also get 5 full-length practice tests, hundreds of sample questions, and all the facts about the current exam. With McGraw-Hill's SAT Subject Test Chemistry, we'll guide you step-by-step through your preparation program and give you the tools you need to succeed. 5 full-length sample tests with complete explanations for every question. 40 top test items to remember on exam day. A step-by-step review of all topics covered on the exam. Teacher-recommended tips and strategies to help you raise your score.

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