

Chapter 12 – The Behavior of Gases

Part 1 – Assignment: Pressure

- (1) 0.568 atm (2) 69100 Pa (3) 15.85 psi
(4) 920. mmHg (5) 62.8 mmHg (6) 69.0 kPa
- (7) There is less air (less atmosphere) at the top of a mountain because there is less gravity holding the air toward the earth. Since there are fewer particles, there are less collisions. Less collisions mean less atmospheric pressure.
- (8) Collisions of the gas particles against the container walls cause pressure. If you increase the speed of the molecules, there are more collisions and there would be more pressure.
- (9) To increase the pressure, you need to increase collisions. (1) Increase the temperature – creates more collisions and more pressure (2) Decrease the volume – creates more collisions and more pressure and (3) Increase the quantity of gas – creates more collisions and more pressure

Part 2 – Assignment: The ABCs of Gases – Aecgadre, Boyle, & Charles

- (1) 1.78atm (2) 16.0L (3) 8.10L
- (4) Pressure and quantity independently affect temperature and volume. Charles Law applies only to volume and temperature changing.
- (5) Gravity pulls the atmospheric gases down toward the center of the earth. Measured in atmospheres.
- (6) Absolute temperature scale = Kelvin scale. 0K = -273°C
- (7a) same (7b) same (7c) same (7d) different
- (8) 17.0g N₂ (9) 584K (10) 1890 mmHg

Part 3 – Notes: Ideal Gas Law and Combined Gas Law

- (1) stays the same (2) 1.5 times what (3) 1930 mL (4) 45.8 L
it was originally
- (5) 7.66×10^{22} molecules CO₂ (6) 4.79×10^{11} g N₂ (7) 253K (8) A
- (9) C (10) C (11a) true (11b) false
(11c) true (11d) false

Part 4 – Assignment: Dalton's Law of Partial Pressure

- (1) 4.14 atm (2) 0.105mol O₂ (3) 3.50atm H₂ and (4) 39.9mL
and 0.0357mol N₂ 0.252 atm N₂
- (5) 19mL (6a) C (6b) C (7) P₂ = 80.6 atm
and 48.1 atm O₂ and
32.5 atm He
- (8a) 29.0 g/mol (8b) 75% N₂ and (9) 0.917 atm
25% O₂

Chapter 12 Chemistry Review Answers Honors

David Baud



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Table of Contents Chapter 12 Chemistry Review Answers Honors

1. Understanding the eBook Chapter 12 Chemistry Review Answers Honors
 - The Rise of Digital Reading Chapter 12 Chemistry Review Answers Honors
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 12 Chemistry Review Answers Honors
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 12 Chemistry Review Answers Honors
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 12 Chemistry Review Answers Honors

- Personalized Recommendations
 - Chapter 12 Chemistry Review Answers Honors User Reviews and Ratings
 - Chapter 12 Chemistry Review Answers Honors and Bestseller Lists
5. Accessing Chapter 12 Chemistry Review Answers Honors Free and Paid eBooks
- Chapter 12 Chemistry Review Answers Honors Public Domain eBooks
 - Chapter 12 Chemistry Review Answers Honors eBook Subscription Services
 - Chapter 12 Chemistry Review Answers Honors Budget-Friendly Options
6. Navigating Chapter 12 Chemistry Review Answers Honors eBook Formats
- ePub, PDF, MOBI, and More
 - Chapter 12 Chemistry Review Answers Honors Compatibility with Devices
 - Chapter 12 Chemistry Review Answers Honors Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chapter 12 Chemistry Review Answers Honors
 - Highlighting and Note-Taking Chapter 12 Chemistry Review Answers Honors
 - Interactive Elements Chapter 12 Chemistry Review Answers Honors
8. Staying Engaged with Chapter 12 Chemistry Review Answers Honors
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 12 Chemistry Review Answers Honors
9. Balancing eBooks and Physical Books Chapter 12 Chemistry Review Answers Honors
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 12 Chemistry Review Answers Honors
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 12 Chemistry Review Answers Honors
- Setting Reading Goals Chapter 12 Chemistry Review Answers Honors
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 12 Chemistry Review Answers Honors

- Fact-Checking eBook Content of Chapter 12 Chemistry Review Answers Honors
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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