

Name _____ Date _____ Class _____

CHAPTER

15

STUDY GUIDE FOR CONTENT MASTERY

Solutions

Section 15.1 What are solutions?

In your notebook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is a(n) **(1)** solution of oxygen gas dissolved in nitrogen gas. The oxygen is an **(2)** solute and nitrogen is the **(3)** solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an **(4)** soluble substance. A substance that does not dissolve is **(5)** insoluble. **(6)** Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are **(7)** miscible. However, if one liquid is insoluble in another, the liquids are **(8)** immiscible.

Read about hydration in aqueous solutions in your notebook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



- Hydration is hydration in which the solvent is water. What is hydration?
Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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CHAPTER

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Section 15.1 continued

- An ionic chloride dissolves in water. What happens to the sodium and chloride ions?
The sodium and chloride ions are separated and surrounded by the water molecules.
- Explain the orientation of the water molecules around the sodium ions and chloride ions.
Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.
- How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?
The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

- Let them say that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your notebook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

- True _____ 14. The overall energy change that occurs when a solution forms is called the heat of solution.
- False _____ 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- False _____ 16. Solvation continues as long as the solvation rate is less than the crystallization rate.
- True _____ 17. In a saturated solution, solvation and crystallization are in equilibrium.
- True _____ 18. Additional water can be dissolved in an unsaturated solution.
- True _____ 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Chapter 15 Study Guide For Content Mastery Answer Key

Michelle Willihnganz, Bruce D. Clayton

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that is both more pressured and more focused than ever on reducing drug errors The increased emphasis includes Chapter 7 Principles of Medication Administration do not confuse icons and an icon to signal ISMP s high alert drugs As the focus on pharmacology on the NCLEX continues to increase chapter ending Get Ready for the NCLEX R Examination sections include Key Points Additional Learning Resources and Review Questions for the NCLEX Examination Updated coverage of the most recent FDA approvals withdrawals and therapeutic uses provide impeccably accurate and current content to ensure safe drug therapy Updated and streamlined coverage of nursing care with a focus on the need to know material helps students focus on the most current need to know content

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Resources in Education ,1975

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