

Acid and Base Worksheet - Answers

- 1) Using your knowledge of the Bronsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair:
- a) $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$
HNO₃ and NO₃⁻ make one pair OH⁻ and H₂O make the other
- b) $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$
CH₃NH₂ and CH₃NH₃⁺ make one pair OH⁻ and H₂O make the other
- c) $\text{OH}^- + \text{HPO}_4^{2-} \rightarrow \text{H}_2\text{O} + \text{PO}_4^{3-}$
HPO₄²⁻ and PO₄³⁻ make one pair OH⁻ and H₂O make the other
- 2) The compound NaOH is a base by all three of the theories we discussed in class. However, each of the three theories describes what a base is in different terms. Use your knowledge of these three theories to describe NaOH as an Arrhenius base, a Bronsted-Lowry base, and a Lewis base.
- **NaOH is an Arrhenius base because it creates OH⁻ ions when placed in water.**
 - **NaOH is a Bronsted-Lowry base because it accepts H⁺ ions from acids.**
 - **NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds.**
- 3) When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base.
- $\text{HCl} + \text{NH}_3 \rightarrow \text{NH}_4\text{Cl}$
NH₃ is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.
HCl is a Lewis acid because it accepts electrons from NH₃ when the H is transferred.
- 4) Write an equation for each of the following reactions and indicate what each reaction produces: (For each of the following oxides, indicate if it is an acid or basic anhydride).
- a. Carbon dioxide and water
 $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$
Acid anhydride
- b. Sodium oxide and water
 $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH(aq)}$
Basic anhydride
- c. Dichlorine trioxide and water
 $\text{Cl}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_3$
Acidic anhydride
- d. Beryllium oxide and water
 $\text{BeO} + \text{H}_2\text{O} \rightarrow \text{Be(OH)}_2$
Basic anhydride

Acid Base Homework Answers

Patricia Ann Mabrouk



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Organic Chemistry Michael B. Smith, 2016-03-09 Based on the premise that many if not most reactions in organic chemistry can be explained by variations of fundamental acid base concepts *Organic Chemistry An Acid Base Approach* provides a framework for understanding the subject that goes beyond mere memorization Using several techniques to develop a relational understanding it helps students fully grasp the essential concepts at the root of organic chemistry This new edition was rewritten largely with the feedback of students in mind and is also based on the author's classroom experiences using the first edition Highlights of the Second Edition Include Reorganized chapters that improve the presentation of material Coverage of new topics such as green chemistry Adding photographs to the lectures to illustrate and emphasize important concepts A downloadable solutions manual The second edition of *Organic Chemistry An Acid Base Approach* constitutes a significant improvement upon a unique introductory technique to organic chemistry The reactions and mechanisms it covers are the most fundamental concepts in organic chemistry that are applied to industry biological chemistry biochemistry molecular biology and pharmacy Using an illustrated conceptual approach rather than presenting sets of principles and theories to memorize it gives students a more concrete understanding of the material *First Year Chemistry Students' Conceptions of Acid/base Chemistry* Sally Diane Rupert, 2001

Foundations of College Chemistry Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years *Foundations of College Chemistry Alternate 14th Edition* has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They'll learn how to apply concepts with the help of worked out examples In addition *Chemistry in Action* features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this

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Table of Contents Acid Base Homework Answers

1. Understanding the eBook Acid Base Homework Answers
 - The Rise of Digital Reading Acid Base Homework Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Homework Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Homework Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Homework Answers
 - Personalized Recommendations
 - Acid Base Homework Answers User Reviews and Ratings
 - Acid Base Homework Answers and Bestseller Lists

5. Accessing Acid Base Homework Answers Free and Paid eBooks
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 - Acid Base Homework Answers Compatibility with Devices
 - Acid Base Homework Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid Base Homework Answers
 - Highlighting and Note-Taking Acid Base Homework Answers
 - Interactive Elements Acid Base Homework Answers
8. Staying Engaged with Acid Base Homework Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Homework Answers
9. Balancing eBooks and Physical Books Acid Base Homework Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Homework Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acid Base Homework Answers
 - Setting Reading Goals Acid Base Homework Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid Base Homework Answers
 - Fact-Checking eBook Content of Acid Base Homework Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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