Practice test Acid-Base Equilibrium

1. For the following reaction of hypochlorous acid (HOCl) with water,

$$HOCl + H_2O \Longleftrightarrow H_3O^+ + OCl^-$$

what would be the effect of adding sodium hypochlorite (NaOCl) to the reaction at equilibrium?

   a. The concentrations of both HOCl and H_3O^+ would increase.
   b. The concentrations of both HOCl and H_3O^+ would decrease.
   c. The concentration of HOCl would increase and the concentration of H_3O^+ would decrease.
   d. The concentration of HOCl would decrease and the concentration of H_3O^+ would increase.
   e. There would be no change because sodium hypochlorite is a salt without any acidic or basic properties.

2. Calculate the molar hydronium ion concentration in a solution containing 0.23 M hypochlorous acid (HOCl), a monoprotic weak acid used in bleach solutions. For HOCl, $K_a = 2.9 \times 10^{-8}$.

3. Given three separate solutions containing equal concentrations of formic acid ($K_a = 1.7 \times 10^{-4}$), phenol ($K_a = 1.3 \times 10^{-10}$), and acetic acid ($K_a = 1.8 \times 10^{-5}$), select the response below that has the acids arranged in order of increasing percent dissociation at equilibrium.
   a. formic < phenol < acetic
   b. formic < acetic < phenol
   c. acetic < formic < phenol
   d. phenol < acetic < formic
   e. No response is correct.

4. From the following choices, select the one that would be the most basic (least acidic).
   a. 0.1 M hydrochloric acid (a strong acid)
   b. 0.1 M acetic acid (a weak acid)
   c. 0.1 M sodium acetate (the salt of a weak acid)
   d. 0.1 M ammonium chloride (the salt of a weak base)
   e. pure water

Acid Base Equilibrium Practice Test

Shasha Hu

Acid Base Equilibrium Practice Test:

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