

## Answer Key

### Chapter 10

#### Practice Problems

- $9.45 \times 10^{24}$  molecules  $\text{CO}_2$
- $3.27 \times 10^{22}$  molecules  $\text{H}_2\text{O}$
- 1.53 mol Fe
- 0.0341 mol sucrose
- $1.08 \times 10^{21}$  formula units  $\text{CuSO}_4$
- 839 g Sb
- 5.53 g Se
- 6.95 mol S
- 12.5 mol He
- 0.0213 mol Ni
- $6.17 \times 10^{21}$  atoms Pt
- $1.88 \times 10^{28}$  atoms S
- 39.7 g Hg
- $6.35 \times 10^{-3}$  g
- $4.36 \times 10^{16}$  Pb ions
- 0.220 mol  $\text{C}_6\text{H}_6$
- 2.856 mol  $\text{KClO}_3$
- 40.2 g  $\text{SnSO}_4$
- $3.40 \times 10^2$  g (340 g expressed to 3 significant figures)
- 13.1 g Zn and 37.5 g  $\text{Cu}(\text{NO}_3)_2$
- 52.93% Al, 47.07% O
- 16.39% Mg, 18.89% N, 64.72% O
- 39.17% O
- 49.30% N
- The percent compositions will be the same because the ratio of carbon atoms to hydrogen atoms is the same, 1:1.
- $\text{CH}_2\text{O}$
- $\text{ZnN}_2\text{O}_6$
- $\text{Al}_4\text{C}_3$
- $\text{C}_3\text{H}_4\text{O}_3$
- $\text{C}_8\text{H}_8\text{N}_2\text{O}_2$
- $\text{B}_3\text{H}_6\text{N}_3$
- $\text{Ag}_2\text{C}_2\text{O}_4$
- $\text{P}_2\text{S}_5$ ; same as empirical formula
- $\text{C}_9\text{H}_{12}\text{N}_6$
- $\text{CeI}_3 \cdot 9\text{H}_2\text{O}$
- $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$
- $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$
- $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$

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