

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from  $K_w$  and  $K_a$  or  $K_b$  and the concentration of a weak acid or base.
3. Calculate the pH of weak acid and base solutions.

## Vocabulary

- |                   |                             |              |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion  | 5. $K_a$                    | 6. $K_b$     |
|                   | 7. weak base                | 8. pH        |

## Key Equations

- |  |   |
|--|---|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$                         | 4. $\text{pOH} = -\log[\text{OH}^-]$                              |

## Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the  $K_a$  and  $K_b$  values listed in Appendix C. Assume that all solutions are aqueous.

1. What products are **1** formed when hydrochloric acid ( $\text{HCl}$ ) reacts with  $\text{H}_2\text{O}$ ? The concentration of these acids produced in 0.100 M  $\text{HCl}$  solution is **2**  $1.0 \times 10^{-7}$  M.
2. Phosphoric acid has a  $K_a$  of  $7.5 \times 10^{-3}$  at 25°C. Calculate the **3**  $[\text{H}_3\text{O}^+]$  concentration in a 0.100 M solution. **4**  $[\text{H}_3\text{O}^+] = 8.7 \times 10^{-2}$  M. **5**  $[\text{H}_3\text{O}^+] = 8.7 \times 10^{-3}$  M. **6**  $[\text{H}_3\text{O}^+] = 8.7 \times 10^{-4}$  M. **7**  $[\text{H}_3\text{O}^+] = 8.7 \times 10^{-5}$  M.
3. **8**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-7}$  M. **9**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-8}$  M. **10**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-9}$  M. **11**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-10}$  M.
4. **12**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-7}$  M. **13**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-8}$  M. **14**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-9}$  M. **15**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-10}$  M.

16. **16**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-7}$  M. **17**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-8}$  M. **18**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-9}$  M. **19**  $[\text{H}_3\text{O}^+] = 1.0 \times 10^{-10}$  M.
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**Philip Bovier Hawk, Olaf Bergeim**



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