

Double Replacement Reactions and Solubility

Post-Lab Questions: Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1)	$\text{Na}_2\text{CO}_3(\text{aq}) + \text{Al}(\text{NO}_3)_3(\text{aq}) \rightarrow 6\text{NaNO}_3(\text{aq}) + \text{Al}(\text{OH})_3(\text{s})$	None
2)	$\text{NaCl}(\text{aq}) + \text{NH}_4\text{NO}_3(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{NH}_4\text{Cl}(\text{aq})$	None
3)	$\text{NaI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{AgI}(\text{s})$	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
4)	$\text{Na}_3\text{PO}_4(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow 6\text{NaNO}_3(\text{aq}) + (\text{Cu}_3(\text{PO}_4)_2(\text{s}))$	$3\text{Cu}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Cu}_3(\text{PO}_4)_2(\text{s})$

Use your observations table to answer the following. *ALL ions must include a symbol AND a charge or it is wrong!*

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are _____

b) Salts of these anions may be described as SOLUBLE except for those with the cations _____

6a) The symbols and charges of the anion that produced the most (>5) precipitates are _____

b) Salts of these anions may be described as INSOLUBLE except for those with the cations _____

7. Although you did not test the bromide ion (Br^-), would you expect most bromide salts (AlBr_3 , BaBr_2 , CaBr_2 , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. _____

8. Explain your answer in #7 _____

9. What cation combined with bromide is an exception to the rule in #7? _____

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

Double Replacement Reaction And Solubility Lab Key

Katrin Zwirglmaier



Double Replacement Reaction And Solubility Lab Key:

Whole-class Inquiry Dennis W. Smithenry, Joan Gallager-Bolos, 2009 In response to requests from science education professionals this is the perfect vehicle for implementing and assessing this concept of whole class inquiry in your classroom This is a must have package for preservice and inservice middle and high school science teachers Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 **CSIR NET Chemical Sciences Question Bank** Mocktime Publication, The Chemical Sciences question bank for CSIR NET is fully aligned with the official syllabus covering Inorganic Chemistry Organic Chemistry Physical Chemistry Analytical Materials etc It offers PDF downloads of past papers with answers and solutions topic wise question banks advanced problem sets in Part C that require analytical reasoning beyond rote learning The exam pattern identical to the other NET subjects has designated number of questions in Parts A B and C as per CSIR HRDG guidelines Using this bank enables candidates to map syllabus to repeated high weight topics practise book based MCQs and past paper items and refine speed and accuracy under timed conditions Alongside the question bank affiliated books provide conceptual notes solved examples and full length mock test series For aspirants targeting fellowships or lectureships in Chemical Sciences this resource becomes central to structured revision and predictive test simulation *Prentice Hall Chemistry*, 2000 *Improving Student Comprehension in Chemistry Laboratories* Tracy Lynn Haroff, 2006 *Essentials of Introductory Chemistry* Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry Introductory Chemistry Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to

the wealth of technology based support on The Chemistry Place™ Special Edition and on The Chemistry of Life CD ROM For college instructors and students

[Advances in Nanostructured Materials for Multifunctional Applications: Synthesis, Characterization and Properties](#) Luz Stella Gomez-Villalba, 2019-01-25 Special topic volume with invited peer reviewed papers only

[Stereoselective Synthesis of Drugs and Natural Products](#) Vasyl Andrushko, Natalia Andrushko, 2013-08-16 Brings together the best tested and proven stereoselective synthetic methods Both the chemical and pharmaceutical industries are increasingly dependent on stereoselective synthetic methods and strategies for the generation of new chiral drugs and natural products that offer specific 3 D structures With the publication of Stereoselective Synthesis of Drugs and Natural Products researchers can turn to this comprehensive two volume work to guide them through all the core methods for the synthesis of chiral drugs and natural products Stereoselective Synthesis of Drugs and Natural Products features contributions from an international team of synthetic chemists and pharmaceutical and natural product researchers These authors have reviewed the tremendous body of literature in the field in order to compile a set of reliable tested and proven methods alongside step by step guidance This practical resource not only explores synthetic methodology but also reaction mechanisms and applications in medicinal chemistry and drug discovery The publication begins with an introductory chapter covering general principles and methodologies nomenclature and strategies of stereoselective synthesis Next it is divided into three parts Part One General Methods and Strategies Part Two Stereoselective Synthesis by Bond Formation including C C bond formation C H bond formation C O bond formation C N bond formation Other C heteroatom formation and other bond formation Part Three Methods of Analysis and Chiral Separation References in every chapter serve as a gateway to the literature in the field With this publication as their guide chemists involved in the stereoselective synthesis of drugs and natural products now have a single expertly edited source for all the methods they need

The Synthesis and Characterization of Nitrides and Other Non-oxide Materials Mark Stuart Bailey, 2005

Energy Research Abstracts, 1993

Multiscale Operational Organic Chemistry John W. Lehman, 2009 This comprehensive laboratory text provides a thorough introduction to all of the significant operations used in the organic lab and includes a large selection of traditional scale and microscale experiments and minilabs Its unique problem solving approach encourages students to think in the laboratory by solving a scientific problem in the process of carrying out each experiment The Second Edition contains a new introductory section Chemistry and the Environment which includes a discussion of the principles of green chemistry Several green experiments have been added and some experiments from the previous editions have been revised to make them greener

Ei Engineering Conference Index: pt. 1. Civil, environmental, and geological engineering, 1985

[Ei Engineering Conference Index](#), 1985

[Government Reports Announcements](#), 1975-10-03

Government Reports Announcements & Index, 1990-05

[The Encyclopedia of Vitamin E](#) Victor R. Preedy, Ronald Ross Watson, 2007 Vitamin E is an important dietary constituent which helps in the defence against cellular damage The process of its absorption from

food and its utilization by the body is an intricate series of reactions It is also used therapeutically in treating numerous diseases and conditions such as skin damage and the prevention of pathological lesions in major organs and has been shown to be an important factor in preventing heart disease and cancer Over 100 chapters from international contributors make this book the most comprehensive reference work in describing both the positive and negative effects and actions of Vitamin E Chapters are divided into subsections which cover nomenclature biochemical physical and chemical aspects of vitamin E related compounds dietary and nutritional influences and effects cocktails anti oxidants mixtures and novel analogues general physiological systems metabolism and metabolic stress brain neurological and optical systems reproductive systems fetus and infant musculo skeletal systems and exercise cardiovascular and pulmonary systems skin hepatic nephrotic and gastrointestinal systems immune and haematological systems and cancer

Papers Presented at the ... Meeting
American Chemical Society. Division of Polymer Chemistry,1997 **Carolina Science and Math** Carolina Biological Supply Company,2003 EPA Publications Bibliography ,1984

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