

# Physical Chemistry 2nd Edition Ball

## Solutions Manual

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## Chapter 7

### *Equilibria for Multiple Component Systems*

- 7.2. With only a single component and phase, if only  $\text{Fe}_2(\text{SO}_4)_3$  were present in our system, we would have  $C - P + 2 = 1 - 1 + 2 = 2$  degrees of freedom.
- 7.4. For a single-component system, there will never be negative degrees of freedom. The maximum number of phases that can exist in equilibrium is three, so the Gibbs phase rule would yield  $1 - 3 + 2 = 0$  degrees of freedom (like might occur at the triple point). For a two-component system, the maximum number of phases that can exist in equilibrium is four. There will never be negative degrees of freedom.
- 7.6. In this reaction, there are only three independent components (the amount of the fourth can be determined from the equilibrium constant) and 3 phases, two solids and a gas mixture. Therefore, there are  $3 - 3 + 2 = 2$  degrees of freedom.
- 7.8. For water:  $n = \frac{pV}{RT} = \frac{(23.76 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm}/\text{mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 6.39 \times 10^{-3}$  mol of water is needed to ensure that there is a liquid and gas phase. This equals 0.115 grams.  
For methanol:
- $$n = \frac{pV}{RT} = \frac{(125.0 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm}/\text{mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 3.36 \times 10^{-2}$$
- mol of methanol is needed to ensure that there is a liquid and gas phase. This equals 1.08 grams.
- 7.10. Since, according to equation 7.11,  $a_i = \frac{p_i}{p_i^*}$ , and since  $p_i^* = 760 \text{ torr}$ , we can calculate  $a_i$ :

# Ball Solution Manual

**Ensheng Dong**



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Student Solutions Manual for Physical Chemistry David Ball,2002-09 The manual consists of complete solutions to all odd end of chapter exercises and problems

**Student Solutions Manual for C** Daniel Reger,2010-01-01 **Student Solutions Manual, Physical Chemistry Second Edition, David W. Ball** Jörg C. Woehl,2015 **Student Study Guide & Selected Solutions Manual** Frank L. H Wolfs,2008 Instructor's Solutions Manual Charles Henry Edwards,1990

**Fundamentals of Physics, Student's Solutions Manual** David Halliday,J. Richard Christman,1997 This popular book incorporates modern approaches to physics It not only tells readers how physics works it shows them Applications have been enhanced to form a bridge between concepts and reasoning

**Calculus, Student Solutions Manual** Deborah Hughes-Hallett,Andrew M. Gleason,William G. McCallum,Daniel E. Flath,David O. Lomen,David Lovelock,Jeff Tecosky-Feldman,Thomas W. Tucker,Joseph Thrash,Karen R. Rhea,Andrew Pasquale,Sheldon P. Gordon,Douglas Quinney,Patti Frazer Lock,1997-11-10 A revision of the best selling innovative Calculus text on the market Functions are presented graphically numerically algebraically and verbally to give readers the benefit of alternate interpretations The text is problem driven with exceptional exercises based on real world applications from engineering physics life sciences and economics Revised edition features new sections on limits and continuity limits l Hopital s Rule and relative growth rates and hyperbolic functions

*Student Solutions Manual* Michael Sullivan,1999-10-29 Includes the solutions to the odd numbered problems in the textbook

*Student Solutions Manual for Calculus: Early Transcendentals Single Variable* Jon Rogawski,2011-06-24 *Student Solutions Manual for Calculus Late Transcendentals Single Variable*

*Instructor's Solutions Manual to Accompany Introductory Chemistry* Daniele Peters,1997

**Student Solutions Manual** Robert Ellis,Denny Gulick,1986

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**Student Solutions Manual and Study Guide to Accompany Physics for Scientists and Engineers** John R. Gordon,Raymond A. Serway,2004 Written by John R Gordon Ralph McGrew and Raymond Serway the two volume manual features detailed solutions to 20 percent of the end of chapter problems from the text This manual also features a list of important equations concepts and answers to selected end of chapter questions

*Instructor's Solutions Manual, Introductory Statistics, 3rd Edition, Weiss/Hassett*

Bernard J. Morzuch,1991

**Instructor's Solutions Manual to Accompany Introductory Statistics, Fifth Edition, Neil A. Weiss** David Ralph Lund,Christine A. Franklin,Bernard J. Morzuch,1999

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Study Guide, Student Solutions Manual Raymond A. Serway,John R. Gordon,1998

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