

## Chapter 10 Review: Chemical Quantities & Calculations

*(aka – mole & volume calculations, % composition, empirical & molecular formulas)*

For the test, be sure that you are familiar, and able to complete, the problems below as well as the problems found on the following worksheets: "Molecular Weight and Mole Calculations," "Percent Comp. & Chemical Formula Calculations" and "Chapter 10 Sample Problems." Check the class website for answers to the review worksheet.

1. The air around you is made up of 20.95% oxygen. How many oxygen molecules are in 75.0 grams of air? The density of air is 1.19 g/L. *assume STP.*
2. How much helium (*grams*) would be needed to inflate a balloon to a volume of 5.50L? *assume STP.*
3. The molecular formula of an antibacterial drug is  $C_{17}H_{15}FN_2O_5$ . How many fluorine atoms are in 150mg of this drug?
4. A molecule is made up of 40.0% carbon, 6.67% hydrogen, and an unknown amount of oxygen. What is the molecular formula of the molecule if its molar mass is 90.0g/mole?
5. What is the empirical formula for a compound made up of 71.72% chlorine, 16.16% oxygen and 12.12% carbon? Calculate the molecular formula for the compound if it has a molar mass of 296.7g/mol.
6. How many molecules are in 8.42 Kilograms of  $C_2H_4$ ?

# 10 Chemical Quantities Guided Practice Problems Answers

**Ian Pickup**



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of water but also contain salt fatty oils and over 1 500 different proteins

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web dec 28 2016 tears also keep your eyes moist which is important for your vision tear glands produce tears and tear ducts carry the tears from the glands to the surface of your eye problems with the tear system can include too many tears too few tears or problems with the tear ducts treatment of the problem depends on the cause

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web as verbs the difference between tears and tear is that tears is third person singular of tear while tear is to rend a solid material by holding or restraining in two places and pulling apart whether intentionally or not to destroy or separate