

**Data Table 2: Heat of Reaction of HCl and Solid NaOH**

(1) Mass of calorimeter bottom with 100.0mL distilled water	105.182g
(2) Mass of dry, empty calorimeter bottom	5.303g
(3) Mass of solution used	
(4) Mass of weighing jar with pellets	71.499g
(5) Mass of weighing jar after pellets were added to the reaction.	70.442g
(6) Mass of Solid NaOH added	
(7) Total mass in calorimeter	
(8) Highest temperature reached during the reaction	27.44 °C
(9) Initial temperature before adding pellets	21.97 °C
(10) Temperature difference ( $\Delta T$ )	
(11) Heat evolved ( $q = m \times C_s \times \Delta T$ )	
(12) Moles of solid NaOH added	
(13) Heat evolved per mole ( $C_{mol}$ )	
(14) Heat of Reaction for NaOH ( $\Delta H_2$ )	

# Additivity Of Heats Of Reaction Lab Answers

**VM Jensen**



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