

# Relative formula mass

## Practice worksheet

4.1 Calculating relative formula mass (M <sub>r</sub> )	
Calculate the relative molecular mass of water	
$M_r(\text{H}_2\text{O})$	$= 2 \times 1 + 1 \times 16$ $= 2 \times 1 + 1 \times 16$ $= 2 + 16$ $= 18$
Calculate the M <sub>r</sub> of MgO	$M_r(\text{MgO})$ $= 1 \times 24 + 1 \times 16$ $= 1 \times 24 + 1 \times 16$ $= 24 + 16$ $= 40$
Calculate the M <sub>r</sub> of CO <sub>2</sub>	$M_r(\text{CO}_2)$ $= 1 \times 12 + 2 \times 16$ $= 1 \times 12 + 2 \times 16$ $= 12 + 32$ $= 44$
Calculate the M <sub>r</sub> of C <sub>2</sub> H <sub>6</sub>	$M_r(\text{C}_2\text{H}_6)$ $= 2 \times 12 + 6 \times 1$ $= 2 \times 12 + 6 \times 1$ $= 24 + 6$ $= 30$
Calculate the M <sub>r</sub> of H <sub>2</sub> SO <sub>4</sub>	$M_r(\text{H}_2\text{SO}_4)$ $= 2 \times 1 + 1 \times 32 + 4 \times 16$ $= 2 \times 1 + 1 \times 32 + 4 \times 16$ $= 2 + 32 + 64$ $= 98$
Calculate the M <sub>r</sub> of Al <sub>2</sub> O <sub>3</sub>	$M_r(\text{Al}_2\text{O}_3)$ $= 2 \times 27 + 3 \times 16$ $= 2 \times 27 + 3 \times 16$ $= 54 + 48$ $= 102$
Calculate the M <sub>r</sub> of K <sub>2</sub> SO <sub>4</sub>	$M_r(\text{K}_2\text{SO}_4)$ $= 2 \times 39 + 1 \times 32 + 4 \times 16$ $= 2 \times 39 + 1 \times 32 + 4 \times 16$ $= 78 + 32 + 64$ $= 174$

4.1 Calculating relative formula mass (M <sub>r</sub> )	
Calculate the M <sub>r</sub> of Fe <sub>2</sub> O <sub>3</sub>	$M_r(\text{Fe}_2\text{O}_3)$ $= 2 \times 56 + 3 \times 16$ $= 2 \times 56 + 3 \times 16$ $= 112 + 48$ $= 160$
Calculate the M <sub>r</sub> of CaCO <sub>3</sub>	$M_r(\text{CaCO}_3)$ $= 1 \times 40 + 1 \times 12 + 3 \times 16$ $= 40 + 12 + 48$ $= 100$
Calculate the M <sub>r</sub> of Ca(OH) <sub>2</sub>	$M_r(\text{Ca(OH)}_2)$ $= 1 \times 40 + 2 \times 16 + 2 \times 1$ $= 40 + 32 + 2$ $= 74$
Calculate the M <sub>r</sub> of C <sub>2</sub> H <sub>5</sub> OH	$M_r(\text{C}_2\text{H}_5\text{OH})$ $= 2 \times 12 + 5 \times 1 + 1 \times 16 + 1 \times 1$ $= 24 + 5 + 16 + 1$ $= 46$
Calculate the M <sub>r</sub> of MgSO <sub>4</sub>	$M_r(\text{MgSO}_4)$ $= 1 \times 24 + 1 \times 32 + 4 \times 16$ $= 24 + 32 + 64$ $= 120$
Calculate the M <sub>r</sub> of ZnCO <sub>3</sub>	$M_r(\text{ZnCO}_3)$ $= 1 \times 65 + 1 \times 12 + 3 \times 16$ $= 65 + 12 + 48$ $= 125$
Calculate the M <sub>r</sub> of Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	$M_r(\text{Al}_2(\text{SO}_4)_3)$ $= 2 \times 27 + 3 \times 32 + 12 \times 16$ $= 54 + 96 + 192$ $= 342$
Calculate the M <sub>r</sub> of H <sub>2</sub> PO <sub>4</sub>	$M_r(\text{H}_2\text{PO}_4)$ $= 2 \times 1 + 1 \times 31 + 4 \times 16$ $= 2 + 31 + 64$ $= 97$
Calculate the M <sub>r</sub> of H <sub>2</sub> CO <sub>3</sub>	$M_r(\text{H}_2\text{CO}_3)$ $= 2 \times 1 + 1 \times 12 + 3 \times 16$ $= 2 + 12 + 48$ $= 62$
Calculate the M <sub>r</sub> of H <sub>2</sub> SO <sub>4</sub>	$M_r(\text{H}_2\text{SO}_4)$ $= 2 \times 1 + 1 \times 32 + 4 \times 16$ $= 2 + 32 + 64$ $= 98$
Calculate the M <sub>r</sub> of H <sub>2</sub> CO <sub>3</sub>	$M_r(\text{H}_2\text{CO}_3)$ $= 2 \times 1 + 1 \times 12 + 3 \times 16$ $= 2 + 12 + 48$ $= 62$
Calculate the M <sub>r</sub> of H <sub>2</sub> CO <sub>3</sub>	$M_r(\text{H}_2\text{CO}_3)$ $= 2 \times 1 + 1 \times 12 + 3 \times 16$ $= 2 + 12 + 48$ $= 62$

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