

STUDY CARD

For CAMPBELL-REECE BIOLOGY Seventh Edition

Exploring Life (Ch. 1)

1. Life exists from the microscopic scale to the global.
2. Living systems are much more than the sum of their parts.
3. Organisms include a great diversity of species.
4. Evolution accounts for life's unity and diversity.
5. Biologists use discovery, vision, and hypothesis-based science to explore life.

The Chemical Context of Life (Ch. 2)

1. Matter consists of chemical elements in pure form and in combinations called compounds.
2. An element's properties depend on its structure.
 - a. Protons and neutrons are in the nucleus.
 - b. Electrons exist at different energy levels in electron shells (or subshells) in the relative shell.
3. Atoms with incomplete valence shells can form chemical bonds with other such atoms by sharing, gaining, or losing electrons.
 - a. In a covalent bond, two atoms share a pair of electrons. Double or triple bonds can also form.
 - b. Electrons transfer forms ions. An ionic bond is the attraction between oppositely charged ions.
 - c. Weak bonds include hydrogen bonds and van der Waals interactions.
4. Chemical reactions make and break chemical bonds as reactants are converted to products.

Water (Ch. 3)

1. Water molecules are polar. Hydrogen bonding between water molecules gives water unique properties that allow life to exist on Earth.
 - a. Cohesion and adhesion allow capillary action, which sustains during water transport in plants.
 - b. Water's high specific heat and high heat of vaporization moderate temperatures on Earth.
 - c. Ice floats because hydrogen bonds lock water into a crystalline lattice that reduces its density.



Figure 3.12 and 3.13

- a. Water is a versatile solvent.
 - a. Hydrophobic substances have an affinity for non-polar, while hydrophilic substances regard water.
1. Dissociation of water molecules affects the pH.
 - a. Acids, which have a pH < 7.
 - b. Basic solutions have a pH > 7.
 - c. Buffers minimize changes in pH.

Carbon and Molecular Diversity (Ch. 4)

1. Organic chemistry is the study of carbon compounds.
2. The four atoms are the most diverse molecules for forming bonds with up to four other atoms.
 - a. Carbon atoms form the skeletons of most organic molecules. They can bond to C, H, O, N, and S.
 - b. Carbons have the same number of atoms and elements but different structures and properties.
 - c. Thus, small groups are the parts of molecules involved in chemical reactions.

3. ATP is an important energy-transforming molecule: $ADP + P_i \rightarrow ATP + H_2O$.

Macromolecules (Ch. 5)

1. Most macromolecules are polymers synthesized from monomers by the dehydration class of reaction.
 - a. Polymers are characterized by monomers for addition of a water molecule (hydrolysis).
 - b. Variation in the linear sequence of its elements results in the diversity of polymers.
 - c. Four major classes of biological macromolecule cells are found in cells.
2. Carbohydrates include sugars, starch, and cellulose (plants) and glycogen polymers (prokaryotes).
 - a. Most monosaccharides are multiples of $C_3H_6O_3$.
 - b. Starch (in plants) and glycogen (in animals) are glucose polymers that store energy.
 - c. Cellulose (in plants) and chitin (in animals) are polysaccharides that provide structural support.
3. Lipids are hydrophobic.
 - a. Fats (glycerol linked to three fatty acid chains) can be saturated or unsaturated; they store energy.
 - b. Phospholipids (glycerol, two fatty acids, and phosphate) have hydrophobic and hydrophilic regions; they form a bilayer in water.
 - c. Steroids (a carbon skeleton with four fused rings) include cholesterol and sex hormones.
4. Proteins have diverse structures and functions.
 - a. A protein consists of one or more polypeptides, each a polymer of amino acids (each a set of C_2).
 - b. A peptide bond joins the carbonyl end of one amino acid with the amino group of the next.



Figure 4.1A

5. Four levels of protein structure lead to the unique shape and function of each protein.
6. Nucleic acids (DNA and RNA) store and pass on hereditary information.



Figure 4.2B

- a. DNA and RNA are polymers of nucleotides.
- b. A nucleotide has three parts: a 5-carbon sugar, a nitrogenous base, and a phosphate group.
- c. DNA has deoxyribose as the sugar; RNA has ribose.
- d. DNA has A, C, G, and T as bases; RNA has A, C, G, and U.
- e. DNA is a double-stranded helix; RNA is single-stranded.
- f. A gene (hereditary unit) is a stretch of DNA that is transcribed into RNA, which is then translated into the amino acid sequence of a polypeptide.

A Tear of the Cell (Ch. 6)

1. All organisms are composed of one or more cells.
2. Cells are bounded by a plasma membrane and contain cytosol, organelles, and in some eukaryotes, one or more chloroplasts, and many ribosomes.

3. Prokaryotic cells usually have a cell wall but lack a nucleus and other membrane-bounded organelles.



Figure 6.1

4. Eukaryotic cells are generally larger than prokaryotic cells and have a nucleus and membrane-bounded organelles.
 - a. An animal cell.

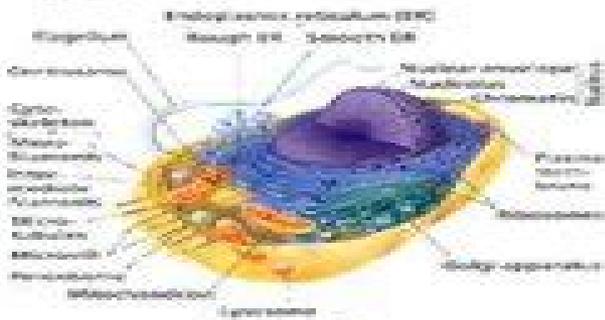


Figure 6.2a

b. A plant cell



Figure 6.2b

5. Within the nucleus, chromosomes made up of chromatin (DNA and proteins) carry the genetic instructions for making polypeptides.
6. Ribosomes are the sites where organelles in which polypeptides are synthesized.
7. The endoplasmic reticulum (ER) is a network of membranes.
 - a. Smooth ER functions in lipid synthesis, carbohydrate metabolism, and detoxification.
 - b. Rough ER has ribosomes that synthesize secretory proteins, which enter the ER lumen. Rough ER also synthesizes membrane proteins and phospholipids.

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