

# TYPES OF CHEMICAL REACTIONS AND THE COPPER CYCLE

## POST-LAB PROBLEMS

1. Balance the following reactions and label the type of reaction.

	Balance Chemical Reaction	Reaction Type
a.	$\text{___ CaCO}_3(s) \rightarrow \text{___ CaO}(s) + \text{___ CO}_2(g)$	
b.	$\text{___ Fe}(s) + \text{___ CuSO}_4(aq) \rightarrow \text{___ FeSO}_4(aq) + \text{___ Cu}(s)$	
c.	$\text{___ Al}(s) + \text{___ O}_2(g) \rightarrow \text{___ Al}_2\text{O}_3(s)$	
d.	$\text{___ Na}_2\text{SO}_4(aq) + \text{___ BaCl}_2(aq) \rightarrow \text{___ NaCl}(aq) + \text{___ BaSO}_4(s)$	
e.	$\text{___ CH}_4(g) + \text{___ O}_2(g) \rightarrow \text{___ CO}_2(g) + \text{___ H}_2\text{O}(l)$	
f.	$\text{___ HCl}(aq) + \text{___ NaOH}(aq) \rightarrow \text{___ NaCl}(aq) + \text{___ H}_2\text{O}(l)$	
g.	$\text{___ Pb}(\text{NO}_3)_2(aq) + \text{___ KI}(aq) \rightarrow \text{___ PbI}_2(s) + \text{___ KNO}_3(aq)$	
h.	$\text{___ H}_2\text{SO}_4(aq) + \text{___ KOH}(aq) \rightarrow \text{___ K}_2\text{SO}_4(aq) + \text{___ H}_2\text{O}(l)$	
i.	$\text{___ C}_5\text{H}_{12}(g) + \text{___ O}_2(g) \rightarrow \text{___ CO}_2(g) + \text{___ H}_2\text{O}(l)$	
j.	$\text{___ CaCl}_2(aq) + \text{___ Na}_2\text{CO}_3(aq) \rightarrow \text{___ CaCO}_3(s) + \text{___ NaCl}(aq)$	
k.	$\text{___ Mg}(s) + \text{___ HCl}(aq) \rightarrow \text{___ MgCl}_2(aq) + \text{___ H}_2(g)$	
l.	$\text{___ Zn}(s) + \text{___ AgNO}_3(aq) \rightarrow \text{___ Zn}(\text{NO}_3)_2(aq) + \text{___ Ag}(s)$	
m.	$\text{___ BaCO}_3(s) \rightarrow \text{___ BaO}(s) + \text{___ CO}_2(g)$	
n.	$\text{___ C}_2\text{H}_4(g) + \text{___ O}_2(g) \rightarrow \text{___ CO}_2(g) + \text{___ H}_2\text{O}(l)$	
o.	$\text{___ NaN}_3(s) \rightarrow \text{___ Na}(s) + \text{___ N}_2(g)$	
p.	$\text{___ Mg}(s) + \text{___ O}_2(g) \rightarrow \text{___ MgO}(s)$	

# Chemical Reactions Copper Post Lab Question Answers

**S Marginson**



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