

Equilibrium

Check Point - 1

1. (a) $K_p = \frac{P_c}{P_a^2 \times P_b} = \frac{1}{P^2} = \text{atm}^{-2}$

2. (a) : (c) is correct because concentrations of H_2 and N_2 decrease while that of NH_3 increases with time and after equilibrium, all of them remain constant.

3. (c) $K = \frac{[\text{CO}_2]^2}{[\text{N}_2\text{O}_4]} = \frac{(2 \times 10^{-3} / 2)^2}{(0.2 / 2)} = \frac{10^{-6}}{10^{-1}} = 10^{-5}$



or $P = 200 \text{ mm Hg}$
 $K = \frac{200 \times 200}{400} = 100$

5. (c) : Partial pressure of 1 atoms (p_1)
 $= \frac{50}{100} \times 10^5 \text{ Pa} = 0.40 \times 10^5 \text{ Pa}$

Partial pressure of I_2 (p_{I_2})
 $= \frac{50}{100} \times 10^5 \text{ Pa} = 0.60 \times 10^5 \text{ Pa}$

$K_p = \frac{p_1^2}{p_{\text{I}_2}} = \frac{(0.4 \times 10^5)^2}{0.60 \times 10^5} = 2.66 \times 10^4$

6. (d) : $\text{A}_2 + 3\text{B}_2 \rightleftharpoons 2\text{AB}_2$; $K_1 = x \text{ mol}^{-2} \text{ L}^2$... (i)

$\text{AB}_2 \rightleftharpoons \frac{1}{2}\text{A}_2 + \frac{3}{2}\text{B}_2$; $K_2 = y \text{ mol L}^{-1}$... (ii)

Equation (ii) is obtained by reversing equation (i) and then dividing by 2.

$\therefore K_2 = \frac{1}{\sqrt{K_1}} \Rightarrow y = \frac{1}{\sqrt{x}} = x^{-1/2}$

7. (b) : For reaction (i),
 $P_1 = P_{\text{CO}_2} = 8 \times 10^{-2}$ (given)

For reaction (ii),

$K_p = \frac{P_{\text{CO}_2}^2}{P_{\text{CO}}} = 2$ (given)

$\therefore 2 = \frac{P_{\text{CO}_2}^2}{8 \times 10^{-2}}$ or, $P_{\text{CO}} = 0.4$

8. (c) : $K_c = \frac{[\text{PCl}_3][\text{Cl}_2]}{[\text{PCl}_5]}$ i.e., $0.5 = \frac{0.2 \times [\text{Cl}_2]}{0.4}$

or, $[\text{Cl}_2] = 1.0 \text{ mol L}^{-1}$

9. (b) : Active mass = molar concentration = n/V

From $PV = nRT$, $\frac{n}{V} = \frac{P}{RT}$

10. (a)

Check Point - 2

- (b)
- (a) : Equilibrium shifts backward.
- (d) : At constant volume, there is no change in concentrations.
- (c) : $n_p = n_r$. Pressure has no effect.
- (c) : As $n_p < n_r$ higher the pressure, greater will be yield of Z. As reaction is exothermic in the forward direction, lower the temperature, greater is the yield of Z.
- (c) : The given reaction is endothermic. So, on increasing temperature, it will shift in the forward direction.
- (c) : $\text{Ice} \rightleftharpoons \text{Water}$
 More vol. less vol.

On applying pressure, equilibrium shifts to the side in which volume is less.

8. (c) : The reaction proceeds from right to left when $Q_c > K_c$, so that Q_c tends to decrease to become equal to K_c .

9. (a) : Change in volume, changes the molar concentration. In (a), $n_p = n_r$. Change in volume will not alter the number of moles in this case.

10. (b) : For a hypothetical reaction



$Q_c = \frac{[\text{C}][\text{D}]}{[\text{A}][\text{B}]}$, Q_c increases

Check Point - 3

1. (b) : $\text{pH} = \frac{\text{p}K_1 + \text{p}K_2}{2} = \frac{2.15 + 7.80}{2} = 4.975$

2. (c) : $\text{pOH} = 14 - \text{pH} = 14 - 8.26 = 5.74$
 $5.74 = \text{p}K_b + \log \frac{[\text{NH}_4^+]}{[\text{NH}_3]} = 4.74 + \log \frac{[\text{NH}_4^+]}{0.1}$

$[\text{NH}_3] = 1 \text{ mol L}^{-1}$

$[\text{NH}_4]_2\text{SO}_4] = \frac{1}{2} = 0.5 \text{ mol L}^{-1}$



$t = 0$	1	0
at eqm.	$1 - \alpha$	2α

Total number of moles at equilibrium = $1 - \alpha + 2\alpha = 1 + \alpha$

$K_p = \frac{P_{\text{NO}_2}^2}{P_{\text{N}_2\text{O}_4}} = \frac{\left(\frac{2\alpha}{1+\alpha}P\right)^2}{\left(\frac{1-\alpha}{1+\alpha}P\right)}$

$K_p = \frac{4\alpha^2 P}{(1+\alpha)(1-\alpha)} = \frac{4\alpha^2 P}{1-\alpha^2}$

$4\alpha^2 P = K_p(1-\alpha^2)$
 $4\alpha^2 P + \alpha^2 K_p = K_p$

Chemical Equilibrium Answer Key

Tao Wei



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Table of Contents Chemical Equilibrium Answer Key

1. Understanding the eBook Chemical Equilibrium Answer Key
 - The Rise of Digital Reading Chemical Equilibrium Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Chemical Equilibrium Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Answer Key
 - Personalized Recommendations
 - Chemical Equilibrium Answer Key User Reviews and Ratings

- Chemical Equilibrium Answer Key and Bestseller Lists
- 5. Accessing Chemical Equilibrium Answer Key Free and Paid eBooks
 - Chemical Equilibrium Answer Key Public Domain eBooks
 - Chemical Equilibrium Answer Key eBook Subscription Services
 - Chemical Equilibrium Answer Key Budget-Friendly Options
- 6. Navigating Chemical Equilibrium Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
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 - Chemical Equilibrium Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Equilibrium Answer Key
 - Highlighting and Note-Taking Chemical Equilibrium Answer Key
 - Interactive Elements Chemical Equilibrium Answer Key
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Equilibrium Answer Key
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Equilibrium Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Equilibrium Answer Key
 - Setting Reading Goals Chemical Equilibrium Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Equilibrium Answer Key
 - Fact-Checking eBook Content of Chemical Equilibrium Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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