

### Intermolecular Force Worksheet Key

1. Identify the strongest intermolecular force present in pure samples of the following substances:



dipole-dipole forces



hydrogen bonds



dipole-dipole forces



dipole-dipole forces



dipole-dipole forces



London dispersion forces

2. Identify the strongest intermolecular force operating in the condensed phases of the following substances. Fully explain how you determined this.

<p>a. Cl<sub>2</sub> London dispersion forces The Cl-Cl bond is nonpolar so the molecule is nonpolar. Non polar molecules have only London dispersion forces operating in the substance.</p>	<p>b. CO Dipole-dipole forces The C-O bond is polar so the molecule is polar. Polar molecules have dipole-dipole forces. They also have London dispersion forces, but dipole-dipole forces are stronger.</p>
<p>c. SO<sub>2</sub> Dipole-dipole forces SO<sub>2</sub> is a bent, polar molecule. The strongest intermolecular force in a polar molecule is the dipole-dipole force</p>	<p>d. CH<sub>2</sub>Cl<sub>2</sub> Dipole-dipole forces The strongest intermolecular force in a polar molecule that cannot form hydrogen bonds is the dipole-dipole force</p>
<p>e. HF Hydrogen bonding forces Molecules that have hydrogen attached to an O, N, or F can form hydrogen bonds. These are the strongest of the intermolecular forces.</p>	<p>g. CH<sub>3</sub>-O-CH<sub>3</sub> Dipole-dipole forces The hydrogen atoms are not bonded to the oxygen, so this molecule cannot form hydrogen bonds. It is polar, so it will have dipole-dipole forces.</p>

3. Based on the intermolecular forces present, predict the relative boiling points of each of the substances below. Arrange each series of substances in order of increasing boiling point. State your reasons for the order you use (identify the forces and explain how they affect the boiling point).

a. dimethyl ether (CH<sub>3</sub>OCH<sub>3</sub>), ethanol (CH<sub>3</sub>CH<sub>2</sub>OH), and propane (CH<sub>3</sub>CH<sub>2</sub>CH<sub>3</sub>)

lowest bp: propane (CH<sub>3</sub>CH<sub>2</sub>CH<sub>3</sub>) < dimethyl ether (CH<sub>3</sub>OCH<sub>3</sub>) < ethanol (CH<sub>3</sub>CH<sub>2</sub>OH) highest bp

Dimethyl ether cannot form hydrogen bonds (no O-H bond), but is polar and has dipole-dipole forces. Ethanol can form hydrogen bonds. propane is nonpolar, so it has only London dispersion forces. The boiling point increases as the strength of the intermolecular forces increase:

London dispersion < dipole-dipole forces < hydrogen bonds

# Chem Activity Intermolecular Forces Key

**Robert R. Redfield**



**Chem Activity Intermolecular Forces Key:**

## **Chem Activity Intermolecular Forces Key** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has been evident than ever. They have the capacity to inspire, provoke, and ignite change. Such may be the essence of the book **Chem Activity Intermolecular Forces Key**, a literary masterpiece that delves deep to the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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