

Annotated periodic table of the elements

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H 1.01 Hydrogen	Main common oxidation states, electronegativity (above-Pauling)																2 He 4.003 Helium
2 Li 6.94 Lithium	3 Be 9.012 Beryllium																
3 Na 22.99 Sodium	4 Mg 24.31 Magnesium																
4 K 39.10 Potassium	5 Ca 40.08 Calcium																
5 Rb 85.47 Rubidium	6 Sr 87.62 Strontium																
6 Cs 132.91 Cesium	7 Ba 137.33 Barium																
7 Fr 223.028 Francium	8 Ra 226.025 Radium																
Atomic number and symbol above, weight and name below		9 Sc 44.96 Scandium	10 Ti 47.87 Titanium	11 V 50.94 Vanadium	12 Cr 51.99 Chromium	13 Mn 54.94 Manganese	14 Fe 55.85 Iron	15 Co 58.93 Cobalt	16 Ni 58.69 Nickel	17 Cu 63.55 Copper	18 Zn 65.38 Zinc	19 Ga 69.72 Gallium	20 Ge 72.61 Germanium	21 As 74.92 Arsenic	22 Se 78.96 Selenium	23 Br 79.90 Bromine	24 Kr 83.80 Krypton
		25 Y 88.91 Yttrium	26 Zr 91.22 Zirconium	27 Nb 92.91 Niobium	28 Mo 95.94 Molybdenum	29 Tc 98.91 Technetium	30 Ru 101.07 Ruthenium	31 Rh 101.07 Rhodium	32 Pd 106.37 Palladium	33 Ag 107.87 Silver	34 Cd 112.41 Cadmium	35 In 114.82 Indium	36 Sn 118.71 Tin	37 Sb 121.76 Antimony	38 Te 127.60 Tellurium	39 I 126.91 Iodine	40 Xe 131.29 Xenon
		57 to 70		71 La 138.91 Lanthanum	72 Ce 140.12 Cerium	73 Pr 140.91 Praseodymium	74 Nd 144.24 Neodymium	75 Pm 144.91 Promethium	76 Sm 150.36 Samarium	77 Eu 151.96 Europium	78 Gd 157.25 Gadolinium	79 Tb 158.93 Terbium	80 Dy 162.50 Dysprosium	81 Ho 164.93 Holmium	82 Er 167.26 Erbium	83 Tm 168.93 Thulium	84 Yb 173.05 Ytterbium
		89 to 102		89 La 138.91 Lanthanum	90 Ce 140.12 Cerium	91 Pr 140.91 Praseodymium	92 Nd 144.24 Neodymium	93 Pm 144.91 Promethium	94 Sm 150.36 Samarium	95 Eu 151.96 Europium	96 Gd 157.25 Gadolinium	97 Tb 158.93 Terbium	98 Dy 162.50 Dysprosium	99 Ho 164.93 Holmium	100 Er 167.26 Erbium	101 Tm 168.93 Thulium	102 Yb 173.05 Ytterbium
		103 Lu 174.96 Lutetium	104 Hf 178.49 Hafnium	105 Ta 180.95 Tantalum	106 W 183.84 Tungsten	107 Re 186.21 Rhenium	108 Os 190.23 Osmium	109 Ir 192.22 Iridium	110 Pt 195.08 Platinum	111 Au 196.97 Gold	112 Hg 200.59 Mercury	113 Tl 204.38 Thallium	114 Pb 207.2 Lead	115 Bi 208.98 Bismuth	116 Po 209 Polonium	117 At 208.98 Astatine	118 Rn 222.018 Radon
		103 Lu 174.96 Lutetium	104 Hf 178.49 Hafnium	105 Ta 180.95 Tantalum	106 W 183.84 Tungsten	107 Re 186.21 Rhenium	108 Os 190.23 Osmium	109 Ir 192.22 Iridium	110 Pt 195.08 Platinum	111 Au 196.97 Gold	112 Hg 200.59 Mercury	113 Tl 204.38 Thallium	114 Pb 207.2 Lead	115 Bi 208.98 Bismuth	116 Po 209 Polonium	117 At 208.98 Astatine	118 Rn 222.018 Radon

- Group names*
- 1 Hydrogen & the alkali metals
 - 2 Alkaline earth metals
 - 3 Lanthanoids
 - 4 Actinoids
 - 5 Boron group
 - 6 Carbon group
 - 7 Nitrogen group
 - 8 Oxygen group
 - 9 Halogens
 - 10 Helium group
 - 11 Crystallogens
 - 12 Pnictogens
 - 13 Chalcogens
 - 14 Halogens
 - 15 Helium group

*Groups 1 to 10 are named after their first members i.e. Group 1 is the Sodium Group

CLASS COUNTS/PARTS

Metals	Nonmetals
Pre-transition	Composite
Active	
Transition	Intermediate
Post-transition	Metalloid
Radioactive	Radioactive
Isotopes (Fr-Po, Ra-Ac)	Isotope gas

89 La 138.91 Lanthanum	90 Ce 140.12 Cerium	91 Pr 140.91 Praseodymium	92 Nd 144.24 Neodymium	93 Pm 144.91 Promethium	94 Sm 150.36 Samarium	95 Eu 151.96 Europium	96 Gd 157.25 Gadolinium	97 Tb 158.93 Terbium	98 Dy 162.50 Dysprosium	99 Ho 164.93 Holmium	100 Er 167.26 Erbium	101 Tm 168.93 Thulium	102 Yb 173.05 Ytterbium
103 Lu 174.96 Lutetium	104 Hf 178.49 Hafnium	105 Ta 180.95 Tantalum	106 W 183.84 Tungsten	107 Re 186.21 Rhenium	108 Os 190.23 Osmium	109 Ir 192.22 Iridium	110 Pt 195.08 Platinum	111 Au 196.97 Gold	112 Hg 200.59 Mercury	113 Tl 204.38 Thallium	114 Pb 207.2 Lead	115 Bi 208.98 Bismuth	116 Po 209 Polonium

Radioactive
Approximate data for other elements is given in red

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; inertlike oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and ionic species such as zincates, stannates, and bismuthates.

Metalloids
Metallic looking brittle, fair electrical conductors. All known in amorphous forms. Moderate ionisation energies (750-1,000 kJ/mol), and electronegativity values (1.8-2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colourless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metalloids to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

Group 1
The starts as Sr+H, and then bifurcates into a La branch (i.e. La-Ac), and a Lu branch (i.e. Lu-Lr). Trends going down Sr+H, and the La branch, resemble those of groups 1-2.

Composite nonmetals
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically composite nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colourless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Bringing It The Periodic Table Answers

Marcel A. Müller



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Academic Computing, 1987 *The Chemical News*, 1928 **Science Examinations ... Reports, Etc**, 1909 *General Chemistry for Colleges* B. Smith Hopkins, 1930

The Pharmaceutical Era, 1891 *The Chemical News and Journal of Industrial Science*, 1928 **Experiments and Exercises in Basic Chemistry** Steven Murov, Brian Stedjee, 2000-02-03 **Chemical News and Journal of Industrial Science**, 1928 *Handbook of Financial Mathematics* Justin Hartley Moore, 1929 **General Chemistry** Hermann Irving Schlesinger, 1930 **Nature** Sir Norman Lockyer, 1900 *The National Provisioner*, 1924 **Chemical News and Journal of Physical Science**, 1919 **Science & Culture**, 1952 *Educational Times*, 1923 **The Century**

Dictionary and Cyclopaedia William Dwight Whitney, Benjamin Eli Smith, 1909

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