

Chemistry SOL Review

by Anne Mooring (Jamestown High School, Williamsburg VA, 2006)

Part 2: Atomic Structure and Periodic Relationships

1. Parts of the Periodic Table
2. Introduction to the SOL Periodic Table
3. Atomic Structure and the Periodic Table
4. Periodic Trends in Atomic Radius, Ionization Energy, and Electronegativity.
5. Isotopes and Average Atomic Mass
6. Historical and Quantum Models of the Atom

A calculator will be helpful, but not necessary for this powerpoint.

This section represents 8/50 of the SOL questions

Periodic Table of the Elements

Legend for Element Information:

- Atomic Number: 14
- Symbol: Si
- Atomic Name: Silicon
- Electron Configuration: $1s^2 2s^2 2p^6 3s^2 3p^2$
- Selected Oxidation States: +4, -4
- Group: 14

Periodic Table Structure:

- Groups:** 1, 2, 3-10 (Transition Elements), 11-12, 13, 14, 15, 16, 17, 18
- Periods:** 1, 2, 3, 4, 5, 6, 7
- Special Series:** Lanthanoid Series (bottom left), Actinoid Series (bottom right)
- Notes:** Most transition element properties are those of the most stable or most common isotope.

Group 1	Group 2	Transition Elements										Group 13	Group 14	Group 15	Group 16	Group 17	Group 18
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21
5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22
6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23
7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24

Lanthanoid Series: Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu

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