

Name \_\_\_\_\_

**MULTIPLE CHOICE.** Choose the one alternative that best completes the statement or answers the question.

- The conjugate base of  $\text{HSO}_4^-$  is  
 A)  $\text{H}_2\text{SO}_4$       B)  $\text{SO}_4^{2-}$       C)  $\text{H}_3\text{SO}_4^+$       D)  $\text{HSO}_4^+$       E)  $\text{OH}^-$
- The conjugate acid of  $\text{HSO}_4^-$  is  
 A)  $\text{H}^+$       B)  $\text{H}_2\text{SO}_4$       C)  $\text{HSO}_4^+$       D)  $\text{SO}_4^{2-}$       E)  $\text{HSO}_3^+$
- The molar concentration of hydronium ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_  
 A) 7.00      B)  $1.0 \times 10^{-7}$       C) 1.00      D)  $1.0 \times 10^{-14}$       E) 0.00
- The molar concentration of hydroxide ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_  
 A) 1.00      B)  $1.0 \times 10^{-14}$       C)  $1.0 \times 10^{-7}$       D) 0.00      E) 7.00
- What is the pH of an aqueous solution at  $25^\circ\text{C}$  in which  $[\text{OH}^-]$  is 0.0025 M?  
 A) +2.60      B) -2.60      C) +11.40      D) -11.40      E) -2.25
- What is the pH of an aqueous solution at  $25^\circ$  that contains  $3.98 \times 10^{-9}$  M hydronium ion?  
 A) 7.000      B) 9.000      C) 8.400      D) 5.600      E) 3.980
- What is the concentration (in M) of hydronium ions in a solution at  $25^\circ\text{C}$  with  $\text{pH} = 4.282$ ?  
 A) 4.28      B)  $1.92 \times 10^{-10}$       C)  $1.66 \times 10^4$       D) 9.71      E)  $5.22 \times 10^{-5}$
- Which solution below has the highest concentration of hydroxide ions?  
 A)  $\text{pH} = 3.21$       B)  $\text{pH} = 9.82$       C)  $\text{pH} = 7.93$       D)  $\text{pH} = 12.59$       E)  $\text{pH} = 7.00$
- What is the pH of a 0.015-M aqueous solution of barium hydroxide?  
 A) 12.18      B) 1.52      C) 12.48      D) 1.82      E) 10.35
- Sodium hydroxide is a strong base. This means that \_\_\_\_\_  
 A) aqueous solutions of  $\text{NaOH}$  contain equal concentrations of  $\text{H}^+$  (aq) and  $\text{OH}^-$  (aq)  
 B)  $\text{NaOH}$  does not dissociate at all when it is dissolved in water  
 C)  $\text{NaOH}$  dissociates completely to  $\text{Na}^+$ (aq) and  $\text{OH}^-$ (aq) when it dissolves in water  
 D)  $\text{NaOH}$  cannot be neutralized by a weak acid  
 E)  $\text{NaOH}$  cannot be neutralized by ordinary means
- Of the following acids, \_\_\_\_\_ is not a strong acid.  
 A)  $\text{HNO}_3$       B)  $\text{HCl}$       C)  $\text{HNO}_2$       D)  $\text{HClO}_4$       E)  $\text{H}_2\text{SO}_4$
- Of the following, \_\_\_\_\_ is a weak acid.  
 A)  $\text{HBr}$       B)  $\text{HF}$       C)  $\text{HClO}_4$       D)  $\text{HCl}$       E)  $\text{HNO}_3$

# 2013 Ap Chemistry Practice Test

**Roman Wölfel**



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## **Table of Contents 2013 Ap Chemistry Practice Test**

1. Understanding the eBook 2013 Ap Chemistry Practice Test
  - The Rise of Digital Reading 2013 Ap Chemistry Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying 2013 Ap Chemistry Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2013 Ap Chemistry Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2013 Ap Chemistry Practice Test
  - Personalized Recommendations
  - 2013 Ap Chemistry Practice Test User Reviews and Ratings
  - 2013 Ap Chemistry Practice Test and Bestseller Lists
5. Accessing 2013 Ap Chemistry Practice Test Free and Paid eBooks

- 2013 Ap Chemistry Practice Test Public Domain eBooks
  - 2013 Ap Chemistry Practice Test eBook Subscription Services
  - 2013 Ap Chemistry Practice Test Budget-Friendly Options
6. Navigating 2013 Ap Chemistry Practice Test eBook Formats
    - ePub, PDF, MOBI, and More
    - 2013 Ap Chemistry Practice Test Compatibility with Devices
    - 2013 Ap Chemistry Practice Test Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of 2013 Ap Chemistry Practice Test
    - Highlighting and Note-Taking 2013 Ap Chemistry Practice Test
    - Interactive Elements 2013 Ap Chemistry Practice Test
  8. Staying Engaged with 2013 Ap Chemistry Practice Test
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers 2013 Ap Chemistry Practice Test
  9. Balancing eBooks and Physical Books 2013 Ap Chemistry Practice Test
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection 2013 Ap Chemistry Practice Test
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine 2013 Ap Chemistry Practice Test
    - Setting Reading Goals 2013 Ap Chemistry Practice Test
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of 2013 Ap Chemistry Practice Test
    - Fact-Checking eBook Content of 2013 Ap Chemistry Practice Test
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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