

Answer **three** questions **only** from this section.

2. (a) Complete the table below.

Reaction	Colour of gas produced	Smell of gas produced	Effect of gas on damp red litmus paper
$\text{HCl} + \text{KMnO}_4$			
$\text{NH}_4\text{Cl} + \text{Ca}(\text{OH})_2$			
$\text{Cu} + \text{HNO}_3$			

[9 marks]

(b) (i) Write the chemical formula of an oxide of an element in period three of the periodic table that is:

- (α) neutral ;
 (β) acidic ;
 (γ) basic ;
 (δ) amphoteric.

(ii) Mention **one** property of diamond that makes it suitable for:

- (α) making jewellery ;
 (β) cutting and drilling.

[6 marks]

(c) (i) Name **two** salts that would decompose on heating to produce oxygen gas.

(ii) Write balanced equations for the decomposition reaction for each of the salts mentioned in 2(c)(i).

[6 marks]

(d) Explain briefly why CCl_4 has a lower boiling point than CBr_4 .

[4 marks]

(e) (i) (α) Draw the structures of all possible isomers of the compound with molecular formula $\text{C}_4\text{H}_9\text{OH}$.(β) Name **each** of the isomers drawn in 3(e)(i) (α).(ii) Name the type of isomerism shown by the isomers drawn in 3(e)(i) (α).(iii) State what will be observed when propan-1-ol is oxidized by acidified $\text{K}_2\text{Cr}_2\text{O}_7$.(iv) Name the product formed when propan-1-ol is dehydrated with concentrated H_2SO_4 .

[12 marks]

Turn over

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