

Chapter 1

Some Basic Concepts of Chemistry

Answer 1.9:

The average atomic mass of chlorine

$$\begin{aligned}
 &= \left[(\text{Fractional abundance of } ^{35}\text{Cl}) (\text{Molar mass of } ^{35}\text{Cl}) \right] \\
 &+ \left[(\text{Fractional abundance of } ^{37}\text{Cl}) (\text{Molar mass of } ^{37}\text{Cl}) \right] \\
 &= \left[\left\{ \left(\frac{75.77}{100} \right) (34.9689\text{u}) \right\} + \left\{ \left(\frac{24.23}{100} \right) (36.9659\text{u}) \right\} \right] \\
 &= 26.4959 + 8.9568 = 35.4527 \text{ u}
 \end{aligned}$$

The average atomic mass of chlorine = 35.4527 u

Question 1.10:

In three moles of ethane (C_2H_6), calculate the following:

1. Number of moles of carbon atoms.
2. Number of moles of hydrogen atoms.
3. Number of molecules of ethane.

Answer 1.10:

1. 1 mole of C_2H_6 contains 2 moles of carbon atoms.
 \therefore Number of moles of carbon atoms in 3 moles of $\text{C}_2\text{H}_6 = 2 \times 3 = 6$
2. 1 mole of C_2H_6 contains 6 moles of hydrogen atoms.
 \therefore Number of moles of carbon atoms in 3 moles of $\text{C}_2\text{H}_6 = 3 \times 6 = 18$
3. 1 mole of C_2H_6 contains 6.023×10^{23} molecules of ethane.
 \therefore Number of molecules in 3 moles of C_2H_6
 $= 3 \times 6.023 \times 10^{23} = 18.069 \times 10^{23}$

Question 1.11:

What is the concentration of sugar ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$) in mol L^{-1} if its 20g are dissolved in enough water to make a final volume up to 2L.

Answer 1.11:

Molarity (M) of a solution is given by,

$$\begin{aligned}
 &= \frac{\text{Number of moles of solute}}{\text{Volume of solution in Litres}} = \frac{\text{Mass of sugar/molar mass of sugar}}{2\text{L}} \\
 &= \frac{20\text{g}/[(12 \times 12) + (1 \times 22) + (11 \times 16)] \text{ g}}{2\text{L}} = \frac{20\text{g}/342 \text{ g}}{2\text{L}} = \frac{0.0585 \text{ mol}}{2\text{L}} = 0.02925 \text{ mol L}^{-1}
 \end{aligned}$$

\therefore Molar concentration of sugar = $0.02925 \text{ mol L}^{-1}$

13 Chemistry Exemplar

Shasha Hu



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Table of Contents 13 Chemistry Exemplar

1. Understanding the eBook 13 Chemistry Exemplar
 - The Rise of Digital Reading 13 Chemistry Exemplar
 - Advantages of eBooks Over Traditional Books
2. Identifying 13 Chemistry Exemplar
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 13 Chemistry Exemplar
 - User-Friendly Interface
4. Exploring eBook Recommendations from 13 Chemistry Exemplar
 - Personalized Recommendations
 - 13 Chemistry Exemplar User Reviews and Ratings
 - 13 Chemistry Exemplar and Bestseller Lists

5. Accessing 13 Chemistry Exemplar Free and Paid eBooks
 - 13 Chemistry Exemplar Public Domain eBooks
 - 13 Chemistry Exemplar eBook Subscription Services
 - 13 Chemistry Exemplar Budget-Friendly Options
6. Navigating 13 Chemistry Exemplar eBook Formats
 - ePub, PDF, MOBI, and More
 - 13 Chemistry Exemplar Compatibility with Devices
 - 13 Chemistry Exemplar Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 13 Chemistry Exemplar
 - Highlighting and Note-Taking 13 Chemistry Exemplar
 - Interactive Elements 13 Chemistry Exemplar
8. Staying Engaged with 13 Chemistry Exemplar
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 13 Chemistry Exemplar
9. Balancing eBooks and Physical Books 13 Chemistry Exemplar
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 13 Chemistry Exemplar
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 13 Chemistry Exemplar
 - Setting Reading Goals 13 Chemistry Exemplar
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 13 Chemistry Exemplar
 - Fact-Checking eBook Content of 13 Chemistry Exemplar
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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