

Name: _____ Date: _____ Period: _____

What happened to
Avogadro when he
got bit by 6.02×10^{23}
mosquitoes?
He got Mol-aria

POGIL- Stoichiometry

How do chemists use balanced chemical equations?

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called **Stoichiometry- (stoi-key-ah-meh-tree)** Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: *How do I convert between different chemical species in a given reaction?*

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole

Question: Given the following equation:



How many moles of H_2 are needed to produce 34.8 moles of NH_3 ?

Answer:

$$\frac{34.8 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

OR

$$34.8 \text{ mol NH}_3 \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Key Questions:

1. What was the given in the example above?
2. Where did the $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$ come from?
3. Double check to see if 52.2 moles of H_2 is the correct answer. SHOW the math below.

Basic Stoichiometry Pogil Activity

RM Cervero



Basic Stoichiometry Pogil Activity:

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