KEY

3. The pH in a 0.25 M solution of the acid HBrO is 4.65 . Using this, calculate the value of Ka for the acid HBrO. (4 marks) pH = 4.65 so $[H_3O^+]$ = antilog (−4.65) = 2.239×10^{-5} M

$HBrO + H_2O \rightleftharpoons H_3O^+ + BrO^-$

[I]	0.25		0	0
[C]	-2.239×10^{-5}		$+2.239\times10^{-5}$	$+2.239\times10^{-5}$
[E]	$0.25-2.239\times10^{-5}$		2.239×10^{-5}	2.239×10^{-5}

4

$$K_a = \frac{[H_3O^+][BrO^-]}{[HBrO]} = \frac{(2.239\times10^{-5})^2}{(0.25-2.239\times10^{-5})} = 2.0\times10^{-9}$$

Answer $K_a = 2.0\times10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) pH = 2.355 $[H_3O^+]$ = antilog (−2.355) = 0.004416 M

$C_6H_5COOH + H_2O \rightleftharpoons H_3O^+ + C_6H_5COO^-$

[I]	C_o		0	0
[C]	−0.004416		+0.004416	+0.004416
[E]	C_o − 0.004416		0.004416	0.004416

4

$$K_a = \frac{[H_3O^+][C_6H_5COO^-]}{[C_6H_5COOH]}$$

$$K_a = \frac{(0.004416)^2}{(C_o - 0.004416)}$$

$$C_o - 0.004416 = \frac{(0.004416)^2}{6.5\times10^{-5}} = 0.299976$$

$$C_o = 0.299976 + 0.004416 = 0.30439 \text{ M}$$

Answer C_o = 0.30 M

5. Find the value of Kb for the oxalate ion ($C_2O_4^{2-}$). (1 mark)

$$K_{b(C_2O_4^{2-})} = \frac{K_w}{K_{a(HC_2O_4^-)}} = \frac{1.00\times10^{-14}}{6.4\times10^{-5}} = 1.56\times10^{-10}$$

1

Answer $K_b = 1.6\times10^{-10}$

6. The value of Kb for the weak base methylamine (CH_3NH_2) is 4.4 x 10^{-4}. Calculate the value of Ka for the acid $CH_3NH_3^+$. (1 mark)

1

$$K_{a(CH_3NH_3^+)} = \frac{K_w}{K_{b(CH_3NH_2)}} = \frac{1.00\times10^{-14}}{4.4\times10^{-4}} = 2.27\times10^{-11}$$

Answer $K_a = 2.3\times10^{-11}$

10

Basic Stoichiometry Phet Lab Answer Key

D Kirk

Basic Stoichiometry Phet Lab Answer Key:

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