

## Chemistry:

### CHAPTER 1

- Charge of electron  $\approx -1.6 \times 10^{-19}$ ; Avogadro's #  $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$ ;  $E = qV$  (energy = charge on e- x electron volts);  $c = \lambda \nu$  ( $c = 3 \times 10^8$ )
- **Heisenberg uncertainty** – impossible to determine momentum and position simultaneously
- **Pauli exclusion** – no electrons have same quantum #'s
- Max e- in a shell  $\approx 2n^2$ , max e- in subshell  $\approx 4l+2$ 
  - Possible values for l  $\approx 0$  to  $n-1$ , possible values for ml  $\approx -l$  to  $+l$
- s can hold 2 e-, p  $\approx 6$ , d  $\approx 10$ , f  $\approx 14$
- **Aufbau principle** (building up) – lowest  $n+l$  will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- **Hund's** – half filled orbitals with parallel spins (bus seats)
- **Paramagnetic** – unpaired parallel spin e-, attracted to magnetic field
  - **Diamagnetic** – paired e-, repelled by magnetic field
  - **Ferromagnetic** – unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

### CHAPTER 2

- **Z eff** increases from left to right, constant up and down
- **Ionic radius** is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form **hydration complexes**
- **Atomic radius** has the opposite trend of all others



### CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- **Coordinate covalent bond** – both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic  $\approx$  EN difference greater than 1.7, polar covalent  $\approx$  0.5 to 1.7, nonpolar  $\approx$  less than 0.5
  - Ionic are stronger than covalent (which are both stronger than IM forces)
    - Van der Waals occurs more with close proximity and large molecules
- **Dipole moment**  $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
  - **Formal charge = valence – nonbonding – ½ bonding**
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

# Section 2 nuclear Chemistry Study Guide Key

**SB Merriam**



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Study Guide with Answers to Selected Problems H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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and a master of midi chlorian manipulation who lived during the century leading up to the invasion of naboo obsessed with eternal life plagueis experimented with ways to cheat death and create new life from the midi

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