Answers to Problems

Introduction to Spectroscopy, 5th edition

Pavia, Lampman, Kriz, Vyvyan

Corrections to Problems are in dark blue

CHAPTER 1

1. (a) 90.50% carbon; 9.50% hydrogen (b) C_4H_5

2. 32.0% carbon; 5.4% hydrogen; 62.8% chlorine; $C_3H_6Cl_2$

3. $C_2H_5NO_2$

4. 180.2 = molecular mass. Molecular formula is $C_6H_8O_6$.

5. Equivalent weight = 52.3

6. (a) 6 (b) 1 (c) 3 (d) 6 (e) 12

7. The index of hydrogen deficiency = 1. There cannot be a triple bond, since the presence of a triple bond would require an index of hydrogen deficiency of at least 2.

8. (a) 59.96% carbon; 5.75% hydrogen; 34.29% oxygen (b) $C_7H_8O_3$
(c) $C_{21}H_{24}O_9$ (d) A maximum of two aromatic (benzenoid) rings

9. (a) $C_8H_8O_2$ (b) $C_8H_{12}N_2$ (c) $C_7H_8N_2O$ (d) $C_8H_{12}O_4$

10. Molecular formula = $C_8H_{10}N_2O_2$; index of hydrogen deficiency = 6

11. Molecular formula = $C_{21}H_{30}O_2$; index of hydrogen deficiency = 7

12. For the hydrolysis product: molecular formula = $C_6H_{12}O_6$; index of hydrogen deficiency = 1

For the original carbohydrate: molecular formula = $C_{12}H_{22}O_{11}$; index of hydrogen deficiency = 2

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