

Reaction	Partial Pressure Equilibrium Constant	Equation
$\frac{1}{2} \text{H}_2 \rightleftharpoons \text{H}$	$K_1 = x_1 \frac{p^{1/2}}{x_4^{1/2}}$	Reaction 1
$\frac{1}{2} \text{O}_2 \rightleftharpoons \text{O}$	$K_2 = x_2 \frac{p^{1/2}}{x_8^{1/2}}$	Reaction 2
$\frac{1}{2} \text{N}_2 \rightleftharpoons \text{N}$	$K_3 = x_3 \frac{p^{1/2}}{x_{11}^{1/2}}$	Reaction 3
$\frac{1}{2} \text{H}_2 + \frac{1}{2} \text{O}_2 \rightleftharpoons \text{OH}$	$K_5 = \frac{x_5}{x_4^{1/2} x_8^{1/2}}$	Reaction 4
$\frac{1}{2} \text{O}_2 + \frac{1}{2} \text{N}_2 \rightleftharpoons \text{NO}$	$K_7 = \frac{x_7}{x_8^{1/2} x_{11}^{1/2}}$	Reaction 5
$\text{H}_2 + \frac{1}{2} \text{O}_2 \rightleftharpoons \text{H}_2\text{O}$	$K_9 = \frac{x_9}{x_4 x_8^{1/2} p^{1/2}}$	Reaction 6
$\text{CO} + \frac{1}{2} \text{O}_2 \rightleftharpoons \text{CO}_2$	$K_{10} = \frac{x_{10}}{x_6 x_8^{1/2} p^{1/2}}$	Reaction 7

The equilibrium constant for each reaction is defined by the empirical expression:

$$K = 10 \left[a_{11} \ln(T) + \frac{a_{12}}{T} + a_{13} + a_{14}T + a_{15}T^2 \right] \quad (3-21)$$

18 3 Calculating Equilibrium Constants

JG Myers



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