



LIMITING REAGENTS 2

1. What mass of calcium hydroxide is formed when 10.0 g of calcium oxide reacts with 10.0 g of water?

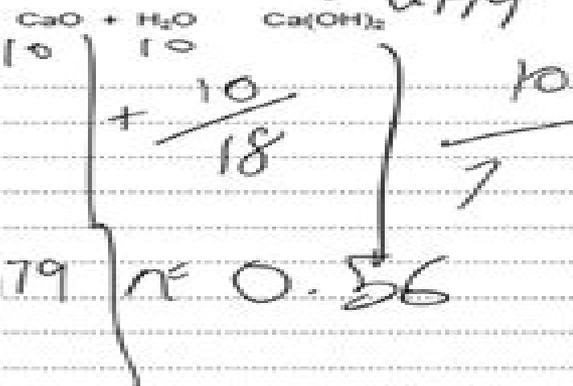
$$n = \frac{m}{Mr}$$

$$m = n \times Mr$$

$$= 74 \times$$

$$n = \frac{10}{74}$$

$$n = 0.179$$



2. What mass of magnesium bromide is formed when 1.00 g of magnesium reacts with 5.00 g of bromine?



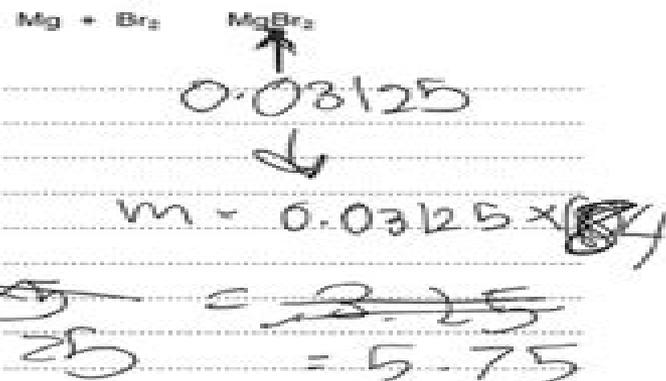
$$n = \frac{m}{Mr}$$

$$= \frac{1}{24}$$

$$= 0.0417$$

$$n = \frac{5}{160}$$

$$= 0.03125$$



3. What mass of copper is formed when 2.00 g of copper(II) oxide reacts with 1.00 g of hydrogen?



$$n = \frac{2}{80} = 0.025$$

$$m = 0.025 \times 64$$

$$n = \frac{1}{2} = 0.5$$

$$= 1.60$$

14 Limiting Reagents Answers

Siddappa N.Byrareddy



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