

3 Scenarios:

1. If the P and T in an all gas system remain constant and the problem is one of volume going to volume...

- **Gay-Lussac's Law** – If the T and P in an all gas system remain constant, the **volumes** of reacting gases are directly related to the **coefficients** of the balanced system.



- Eliminates having to find moles in the problem, thus...
- Eliminates having to use molar volume or $PV = nRT$ in the problem

Example: What volume of hydrogen gas would combine with 42 L. of oxygen gas to form steam at a constant 750 °C and atmospheric pressure?

2. If conditions are at STP but there are grams involved...

- **Molar Volume** – Convert the grams to moles, use the balanced equation, and then use molar volume to find the volume of gas produced.

Example: Calculate the liters of CO₂ measured at 0°C and 1 atm that could be obtained by allowing 45.0 g of calcium carbonate to react with an excess of hydrochloric acid (The molar mass of calcium carbonate is 100. g/mol).

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E Durkheim



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