

2015

Subject : Chemistry (Honours)

Paper : V

Time: 2 Hours

Full Marks: 50

*The figures in the margin indicate full marks.  
Candidates are required to give their answers in their own words  
as far as practicable.*

1. Answer any four questions:

2×4=8

- Depict the ambidentate and flexidentate behaviours of thiocyanate anion.
- $\text{SiCl}_4$  has lower boiling point than  $\text{CCl}_4$ . Explain.
- Cite and draw one example each of  $3c-2e$  and  $2c-2e$  bonds.
- Briefly comment on the compound  $\text{K}_3\text{C}_{60}$ .
- Write down the structural formulae of sodium pyrophosphate and sodium metaphosphate.
- What is meant by chalcogenide? Give example.
- 'NaOH is base but ClOH is acid' - Give reason.

2. Answer any three questions:

6×3=18

- What are inner metallic complexes? Give example. 2
  - Chromium(III) chloride forms three different hydrates of the same mole ratio  $\text{Cr:Cl:H}_2\text{O}=1:3:6$ . The violet form does not lose water over conc.  $\text{H}_2\text{SO}_4$  and gives 3 equivalent of  $\text{AgCl}$  on treatment with  $\text{AgNO}_3$ . The other two forms are green and lose 1 and 2 moles of  $\text{H}_2\text{O}$  over conc.  $\text{H}_2\text{SO}_4$  and give 2 and 1 equivalent of  $\text{AgCl}$ , respectively with  $\text{AgNO}_3$ . Write down the structures of the three isomeric coordination compounds and explain the isomerism involved. 4
- Define and exemplify chelate complex. 3
  - 'The stability of  $[\text{Ni}(\text{en})_3]^{2+}$  (en=ethylenediamine) is much greater than that of  $[\text{Ni}(\text{NH}_3)_6]^{2+}$ , although both contain six Ni-N bonds.' - Explain. 3
- Apply radius ratio principle to explain the difference between crystal structures of NaCl and CsCl. 2
  - 'MgO is a stable ionic solid though large amount of energy is required to form Mg to  $\text{Mg}^{2+}$  and to dissociate  $\text{O}_2$  into atomic oxygen and thereafter to form  $\text{O}^{2-}$  ion.' - Account for the fact. 2

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