

## 19.2 Hydrogen Ions and Acidity

Water molecules can   1   to form hydrogen ions ( $\text{H}^+$ ) and hydroxide ions ( $\text{OH}^-$ ). The concentrations of these ions in pure water at  $25^\circ\text{C}$  are both equal to   2   mol/L.

The pH scale, which has a range from   3  , is used to denote the   4   concentration of a solution. On this scale, 0 is strongly   5  , 14 is strongly   6  , and 7 is   7  . Pure water at  $25^\circ\text{C}$  has a pH of   8  .

The   9   constant for water has a value of  $1.0 \times 10^{-14}$ .

Thus, the product of the concentrations of   10   ions and   11   ions in aqueous solution will always equal  $1.0 \times 10^{-14}$ .

# 19 2 Hydrogen Ions And Acidity Answer

**Ying Liu**



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Carbonic Acid Compounds and Hydrogen Ion Activities in Blood and Salt Solutions Erik Johan Warburg,1922

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