

STOICHIOMETRY

Practice Problems

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SECTION 12.2 CHEMICAL CALCULATIONS

1. A compound contains 40.0% of a single element combined with the element hydrogen.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

2. A compound contains 40.0% of a single element combined with the element hydrogen.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

3. A compound contains 40.0% of a single element combined with the element hydrogen.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

4. A compound contains 40.0% of a single element combined with the element hydrogen.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

5. A compound contains 40.0% of a single element combined with the element hydrogen.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

6. A compound contains 40.0% of a single element combined with the element hydrogen.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

7. A compound contains 40.0% of a single element combined with the element hydrogen.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

8. A mixture of three gases is contained in a 10.0-L container at a pressure of 1.00 atm and a temperature of 273 K. The mixture contains 1.00 mol of gas A, 2.00 mol of gas B, and 3.00 mol of gas C. Calculate the partial pressure of each gas in the mixture.

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

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A Gutmann



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