

**CLASS VIII**  
**SUBJECT: CHEMISTRY**  
**CHAPTER: 6 – 'IONIC BONDING' & CHAPTER: 7 – 'COVALENT & METALLIC BONDING'**  
**PRACTICE QUESTION-ANSWER**

**QUESTION # 1**

write down the formula for the following ionic compounds:

(i) Sodium chloride, (ii) Magnesium oxide, (iii) Aluminium nitride, (iv) Magnesium hydroxide, (v) Iron (III) nitrate, (vi) Ammonium sulphate, (vii) Barium Phosphate

**ANSWER:**

(i) Compound – Sodium chloride

Ions present – Sodium ion & Chloride ion

Formula of ions –  $\text{Na}^+$  &  $\text{Cl}^-$



∴ Formula of compound –  $\text{NaCl}$

(ii) Compound – Magnesium oxide

Ions present – Magnesium ion & Oxide ion

Formula of ions –  $\text{Mg}^{2+}$  &  $\text{O}^{2-}$



∴ Formula of compound –  $\text{MgO}$

(iii) Aluminium nitride

Ions present – Aluminium ion & nitride ion

Formula of ions –  $\text{Al}^{3+}$  &  $\text{N}^{3-}$



∴ Formula of compound –  $\text{AlN}$

# Answers To Section 6 Ionic Bonding Chapter 6

**Michael Rajnik**



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