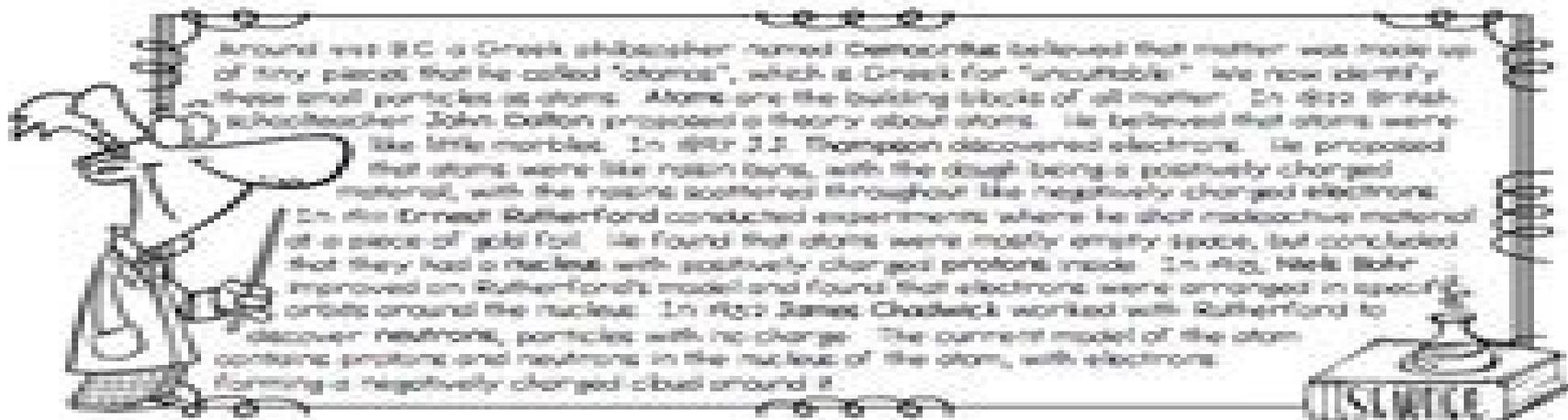


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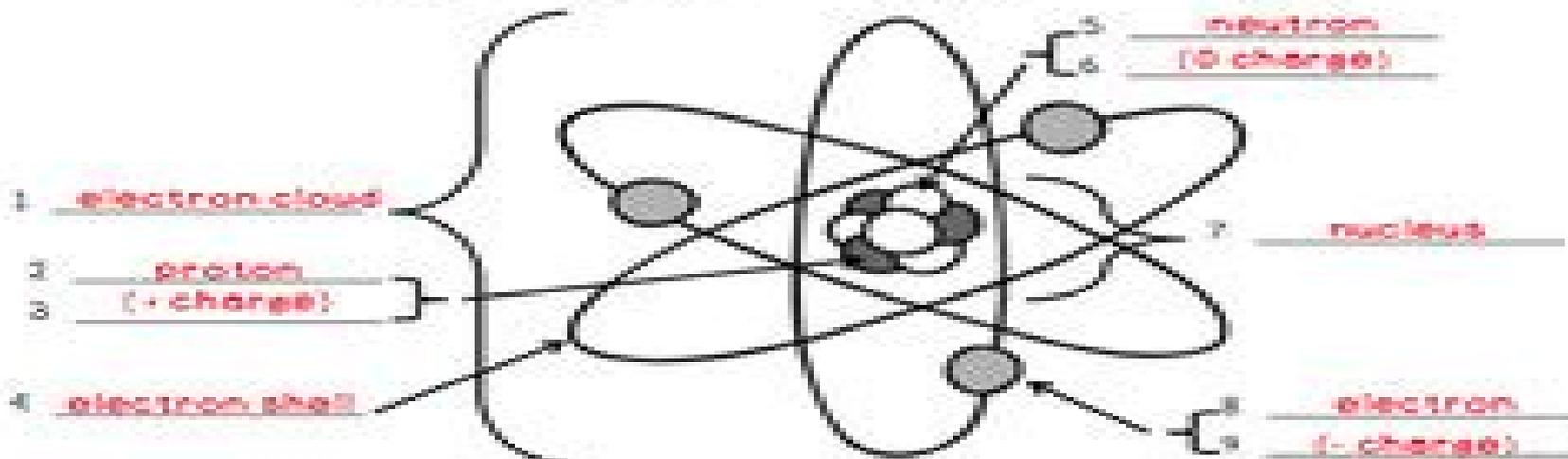


Around 400 B.C. a Greek philosopher named Democritus believed that matter was made up of tiny pieces that he called "atoms", which is Greek for "uncuttable". We now identify these small particles as atoms. Atoms are the building blocks of all matter. In 1808 British schoolteacher John Dalton proposed a theory about atoms. He believed that atoms were like little marbles. In 1897 J.J. Thompson discovered electrons. He proposed that atoms were like mush balls, with the dough being a positively charged material, with the raisins scattered throughout like negatively charged electrons. In 1911 Ernest Rutherford conducted experiments where he shot radioactive material at a piece of gold foil. He found that atoms were mostly empty space, but concluded that they had a nucleus with positively charged protons inside. In 1913, Niels Bohr improved on Rutherford's model and found that electrons were arranged in specific orbits around the nucleus. In 1932 James Chadwick worked with Rutherford to discover neutrons, particles with no charge. The current model of the atom contains protons and neutrons in the nucleus of the atom, with electrons forming a negatively charged cloud around it.

Label It

Correctly label the parts of the atom using the words and phrases below.

- Electron
(Proton)
- Electron cloud
(+ charge)
- Electron shell
(- charge)
- Neutron
(+ charge)
- Nucleus



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Table of Contents Answers To Introduction To Atoms Worksheet

1. Understanding the eBook Answers To Introduction To Atoms Worksheet
 - The Rise of Digital Reading Answers To Introduction To Atoms Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Introduction To Atoms Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To Introduction To Atoms Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Introduction To Atoms Worksheet

- Personalized Recommendations
 - Answers To Introduction To Atoms Worksheet User Reviews and Ratings
 - Answers To Introduction To Atoms Worksheet and Bestseller Lists
5. Accessing Answers To Introduction To Atoms Worksheet Free and Paid eBooks
 - Answers To Introduction To Atoms Worksheet Public Domain eBooks
 - Answers To Introduction To Atoms Worksheet eBook Subscription Services
 - Answers To Introduction To Atoms Worksheet Budget-Friendly Options
 6. Navigating Answers To Introduction To Atoms Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers To Introduction To Atoms Worksheet Compatibility with Devices
 - Answers To Introduction To Atoms Worksheet Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To Introduction To Atoms Worksheet
 - Highlighting and Note-Taking Answers To Introduction To Atoms Worksheet
 - Interactive Elements Answers To Introduction To Atoms Worksheet
 8. Staying Engaged with Answers To Introduction To Atoms Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To Introduction To Atoms Worksheet
 9. Balancing eBooks and Physical Books Answers To Introduction To Atoms Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To Introduction To Atoms Worksheet
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answers To Introduction To Atoms Worksheet
 - Setting Reading Goals Answers To Introduction To Atoms Worksheet
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answers To Introduction To Atoms Worksheet

- Fact-Checking eBook Content of Answers To Introduction To Atoms Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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